



Lab Assistant

# Classifying Chemical Reactions

Directions: Write the type of chemical reaction in the space provided next to each equation. Name the product or the products of the reaction at the end.

## Column I

1. \_\_\_\_\_ 1. A precipitate forms in a gas liquid collection over water (combustion, decomposition, precipitation)
2. \_\_\_\_\_ 2. Two reactants exchange products and/or anions/cations
3. \_\_\_\_\_ 3. A hydrocarbon explodes, usually in a combustion
4. \_\_\_\_\_ 4. One element breaks down into more than one element
5. \_\_\_\_\_ 5. A type of synthesis reaction that produces one molecule

Directions: Write the correct type of chemical reaction in the space provided.

combustion \_\_\_\_\_ 6.  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

decomposition \_\_\_\_\_ 7.  $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$

synthesis

single displacement

\_\_\_\_\_ 8.  $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$

double displacement reaction \_\_\_\_\_ 9.  $\text{NaCl} + \text{KBr} \rightarrow \text{NaBr} + \text{KCl}$

synthesis reaction

\_\_\_\_\_ 10.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

\_\_\_\_\_ 11.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

decomposition

single displacement reaction

\_\_\_\_\_ 12.  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

redox reaction

double displacement reaction

\_\_\_\_\_ 13.  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

decomposition

double displacement reaction

\_\_\_\_\_ 14.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

double displacement reaction \_\_\_\_\_ 15.  $\text{NaCl} + \text{KBr} \rightarrow \text{NaBr} + \text{KCl}$

double displacement reaction

## Section 2 Reinforcement Classifying Chemical Reactions

**Stuart M. Lee**



## Section 2 Reinforcement Classifying Chemical Reactions:

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