

Section 1 Describing Chemical Reactions

A **chemical reaction** is the process by which one or more substances are changed into one or more different substances.

In any chemical reaction, the **original substances** are known as the **reactants**, and the **resulting substances** are known as the **products**.

According to the **law of conservation of mass**, the **total mass of reactants** must equal the **total mass of products** for any given chemical reaction.

Chemical reactions are described by **chemical equations**.

A **chemical equation** represents, with symbols and formulas, the **identities and relative molecular or molar amounts** of the **reactants and products** in a chemical reaction.

Indications of a Chemical Reaction

- **Evolution of energy as heat and light** is strong evidence that a chemical reaction has taken place.
Example, the decomposition of ammonium dichromate is accompanied by the evolution of energy as heat and light.
chemical reaction occurs between natural gas and oxygen if you burn gas for cooking in your house.
- **Production of a gas** the evolution of gas bubbles when two substances are mixed is often evidence of a chemical reaction.
Example, bubbles of carbon dioxide gas form immediately when baking soda is mixed with vinegar.
- **Formation of a precipitate**
a **precipitate** is a solid that is produced as a result of a chemical reaction in solution and that separates from the solution.
- **Color change** is often an indication of a chemical reaction.

Characteristics of Chemical Equations

The following requirements will aid you in writing and reading chemical equations correctly.

1. The equation must represent known facts.
 2. The equation must contain the correct formulas for the reactants and products.
 3. The law of conservation of mass must be satisfied.
- A **coefficient** is a small whole number that appears in front of a formula in a chemical equation.

Section 1 1 Describing Chemical Reactions Answer Key

Peter Baofu



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Numerical Solution of Time-Dependent Advection-Diffusion-Reaction Equations Willem Hundsdorfer, Jan G. Verwer, 2013-04-17 This book deals with numerical methods for solving partial differential equations PDEs coupling advection diffusion and reaction terms with a focus on time dependency A combined treatment is presented of methods for hyperbolic problems thereby emphasizing the one way wave equation methods for parabolic problems and methods for stiff and non stiff ordinary differential equations ODEs With regard to time dependency we have attempted to present the algorithms and the discussion of their properties for the three different types of differential equations in a unified way by using semi discretizations i.e the method of lines whereby the PDE is transformed into an ODE by a suitable spatial discretization In addition for hyperbolic problems we also discuss discretizations that use information based on characteristics Due to this combination of methods this book differs substantially from more specialized textbooks that deal exclusively with numerical methods for either PDEs or ODEs We treat integration methods suitable for both classes of problems This combined treatment offers a clear advantage On the one hand in the field of numerical ODEs highly valuable methods and results exist which are of practical use for solving time dependent PDEs something which is often not fully exploited by numerical PDE researchers Although many problems can be solved by Euler's method or the Crank Nicolson method better alternatives are often available which can significantly reduce the computational effort needed to solve practical problems

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