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# USER MANUAL

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the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c and o by adding h<sup>+</sup> or h<sub>2</sub>O

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni<sup>2+</sup> → ni<sup>3+</sup> + e<sup>-</sup> now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balance each half reaction in basic solution 4 5 6 7 cr<sup>2+</sup> + 7h<sup>+</sup> → cr<sup>3+</sup> + 3h<sub>2</sub> no so<sub>4</sub><sup>2-</sup> + mno<sub>2</sub> + cr<sup>3+</sup> + no<sub>3</sub><sup>-</sup> + so<sub>2</sub> + mn<sub>2</sub>o<sub>3</sub> balance each redox reaction in acid solution using the half reaction method 8 h<sub>2</sub>o<sub>2</sub> + cr<sup>2+</sup> + 2h<sup>+</sup> → teo<sub>2</sub> + 2n<sub>2</sub>o<sub>4</sub> + o<sub>2</sub> + cr<sup>3+</sup> te 10 reo<sub>4</sub><sup>-</sup> + 11 pbo<sub>2</sub> + io<sub>2</sub> + pb<sub>2</sub> 12 as no<sub>3</sub><sup>-</sup> + io<sub>3</sub><sup>-</sup> + re + io<sub>3</sub><sup>-</sup> + h<sub>2</sub>aso<sub>4</sub>

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