

Name/Date \_\_\_\_\_

Teacher/Grade \_\_\_\_\_

Score: Reading Engagement \_\_\_\_/8  
Independent Range: 6-7Oral Reading Fluency \_\_\_\_/16  
11-14Comprehension \_\_\_\_/28  
19-25

Book Selection

Text selected by:

☐ teacher☐ student**1. READING ENGAGEMENT**

(If the student has recently answered these questions, skip this section.)

E Tell me about one of your favorite books. \_\_\_\_\_

E Would you rather read ☐ alone, ☐ with a buddy, or ☐ with a group?

Why? \_\_\_\_\_

E When do you read with at home? \_\_\_\_\_

**2. ORAL READING FLUENCY****INTRODUCTION AND PREVIEW**

E In this story, Robert's New Friend, Robert isn't sure he liked his new baby sister, Maria. Maria and Papa are busy taking care of Maria. Look at the pictures, and tell me what is happening in this story.

Note the student's use of connecting words (e.g., and, then, but) and vocabulary relevant to the text. You may use general prompts, such as "Now what is happening?" or "Turn the page," but do not ask specific questions. Tally the number of times you prompt.

**RECORD OF ORAL READING**

Record the student's oral reading behaviors on the Record of Oral Reading below and on the following page.

E Robert's New Friend. Now, read to see how Robert becomes friends with Maria.

**Page 2**

Robert had a new baby sister.  
She was very little. Her name  
was Maria.

**Page 3**

Robert wasn't sure he liked Maria.  
She cried a lot.

# Roberts New Friend Dra Observe Guide

**Katrin Zwirglmaier**



**Roberts New Friend Dra Observe Guide:**

**El-Hi Textbooks and Serials in Print** ,2006      **El-Hi Textbooks & Serials in Print, 2005** ,2005      *A Study Guide for Robert Louis Stevenson's Dr. Jekyll and Mr. Hyde* Gale, Cengage Learning,2015-09-15 A Study Guide for Robert Louis Stevenson s Dr Jekyll and Mr Hyde excerpted from Gale s acclaimed Novels for Students This concise study guide includes plot summary character analysis author biography study questions historical context suggestions for further reading and much more For any literature project trust Novels for Students for all of your research needs      **Bradshaw's railway almanack, directory, shareholders' guide, and manual** George Bradshaw,1848      **Jeub's Complete Guide to Speech and Debate: For use in home schools, Christian schools, forensics clubs, and individual study** Chris Jeub,2001

**Handbook to London as it is. New edition revised of the work originally written by Peter Cunningham** John Murray (Firm),1867      The New England Historical and Genealogical Register ,1849 Beginning in 1924 Proceedings are incorporated into the Apr number      The Athenæum ,1858      Cult Watch John Ankerberg,John Weldon,1991-01-01 A compilation of the entire Facts On series Cult Watch gives you facts about the major cults and movements in a format you can use in counseling or witnessing      **The Athenaeum** ,1911

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type of chemical reaction for this reaction we have a neutralization reaction balancing strategies here we have a neutralization reaction

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web the balanced chemical equation for this reaction is  $\text{mg} + 2\text{hno}_3 \text{aq} \rightarrow \text{mg}(\text{no}_3)_2 \text{aq} + \text{h}_2 \text{g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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and two moles of nitric acid  $\text{HNO}_3$  react to form one mole of magnesium nitrate  $\text{Mg}(\text{NO}_3)_2$  one mole of water  $\text{H}_2\text{O}$  and one mole of carbon dioxide  $\text{CO}_2$  show chemical structure image reactants

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web the balanced net ionic equation for the reaction of magnesium carbonate with nitric acid is  $\text{MgCO}_3 + 2\text{H}^+ + 2\text{NO}_3^- \rightarrow \text{Mg}^{2+} + \text{CO}_2 + \text{H}_2\text{O} + 2\text{NO}_3^-$  b  $\text{Mg}^{2+} + \text{CO}_3^{2-} + 2\text{H}^+ + 2\text{NO}_3^- \rightarrow \text{Mg}^{2+} + \text{CO}_2 + \text{H}_2\text{O} + 2\text{NO}_3^-$  mg write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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web magnesium oxide nitric acid gives rise to a magnesium phosphate hydrogen balanced equation on may 27 2019 how do you balance magnesium and nitric acid you ll need to count all of atoms on both directions of the chemical equation to balance  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$

**magnesium and nitric acid balanced equation** - Oct 01 2022

web magnesium and nitric acid balanced equation general chemistry oct 29 2021 the eleventh edition was carefully reviewed with an eye toward strengthening the content available in owl2 end of chapter questions and updating the presentation nomenclature changes and the

*how to balance  $\text{HNO}_3 + \text{Mg} \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$*  breslyn org - Sep 12 2023

web how to balance  $\text{HNO}_3 + \text{Mg} \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$  word equation nitric acid magnesium magnesium nitrate hydrogen gas type of chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement reaction the mg replaces the h in the  $\text{HNO}_3$

*how to balance  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$*  magnesium youtube - May 08 2023

web nov 20 2020 in this video we ll balance the equation  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$  and provide the correct coefficients for each compound to balance  $\text{Mg} + \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$  you ll need to be

**lecture 13 acid base reactions widener university** - Dec 03 2022

web magnesium hydroxide nitric acid balanced total equation  $\text{Mg}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$  total ionic equation  $\text{Mg}^{2+} + 2\text{OH}^- + 2\text{H}^+ + 2\text{NO}_3^- \rightarrow \text{Mg}^{2+} + 2\text{NO}_3^- + 2\text{H}_2\text{O}$  net ionic equation  $2\text{OH}^- + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O}$  sulfuric acid and lithium hydroxide

balanced total equation  $\text{H}_2\text{SO}_4 + 2\text{LiOH} \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$