

# Problem Solving: Obtaining Empirical Formula from Experimental Data

## PROCEDURE FOR

### Obtaining an Empirical Formula from Experimental Data

1. Write down (or calculate) in grams the masses of each element present in a sample of the compound. If you are given mass percent composition, assume a 100-g sample and calculate the masses of each element from the given percentages.

2. Convert each of the masses in Step 1 to moles by using the appropriate molar mass for each element as a conversion factor.

3. Write down a preliminary formula for the compound using the number of moles of each element from Step 2 as subscripts.

4. Divide all the subscripts in the formula by the smallest subscript.

5. If the subscripts are not whole numbers, multiply all the subscripts by a small whole number (see table) to get whole-number subscripts.

Fractional Subscript	Multiply by This
0.50	2
0.25	4
0.33	3
0.40	5
0.20	2
0.60	5
0.75	4
0.80	5

## EXAMPLE 5.15

### Obtaining an Empirical Formula from Experimental Data

A laboratory analysis of aspirin determined the following mass percent composition:  
C 60.00%  
H 4.48%  
O 35.52%  
Find the empirical formula.

**GIVEN:** In a 100-g sample, 60.00 g C, 4.48 g H, 35.52 g O.

**Find:** empirical formula.

$$60.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 5.00 \text{ mol C}$$

$$4.48 \text{ g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 4.44 \text{ mol H}$$

$$35.52 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 2.22 \text{ mol O}$$



The correct empirical formula is  $\text{C}_9\text{H}_8\text{O}_4$ .

### FOR PRACTICE 5.15

Bupropion has the following mass percent composition:  
C 76.06% H 6.80% O 17.13%  
What is the empirical formula of Bupropion?

# Problem Solving Empirical Formula Answer Sheet

**E Durkheim**



## **Problem Solving Empirical Formula Answer Sheet:**

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