

## Chapter 1

### *Gases and the Zeroth Law of Thermodynamics*

- 1.2.** A system is any part of the universe under observation. The “surroundings” includes everything else in the universe. Consider a solution calorimeter in which two aqueous solutions are mixed and temperature changes are recorded. In this case the solutes (reactants and products) would be considered the “system”. The water, calorimeter, and rest of the lab and world would be the “surroundings”.
- 1.4.** (a)  $12.56 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 12,560 \text{ cm}^3 = 1.256 \times 10^4 \text{ cm}^3$   
(b)  $45^\circ\text{C} + 273.15 = 318 \text{ K}$   
(c)  $1.055 \text{ atm} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} \times \frac{100,000 \text{ Pa}}{1 \text{ bar}} = 1.069 \times 10^5 \text{ Pa}$   
(d)  $1233 \text{ mmHg} \times \frac{1 \text{ torr}}{1 \text{ mmHg}} \times \frac{1 \text{ atm}}{760 \text{ torr}} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} = 1.644 \text{ bar}$   
(e)  $125 \text{ mL} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 125 \text{ cm}^3$   
(f)  $4.2 \text{ K} - 273.15 = -269.0^\circ\text{C}$   
(g)  $25750 \text{ Pa} \times \frac{1 \text{ bar}}{100,000 \text{ Pa}} = 0.2575 \text{ bar}$
- 1.6.**  $p_{\text{tot}} = p_{\text{atmosph}} + \rho gh$ , where  $\rho gh$  correspond to the pressure exerted by the liquid  
 $p_{\text{tot}} - p_{\text{atmosph}} = \rho gh = (1.0 \times 10^3 \text{ kg/m}^3)(0.23 \text{ m})(9.80 \text{ m/s}^2) = 2254 \text{ N/m}^2 = 2.3 \times 10^3 \text{ Pa}$
- 1.8.** In terms of the zeroth law of thermodynamics, heat will flow from the (hot) burner or flame on the stove into the (cold) water, which gets hotter. Then heat will move from the hot water into the (colder) egg.
- 1.10.** For this sample of gas under these conditions,  $F(T) = 2.97 \text{ L} \times 0.0553 \text{ atm} = 0.164 \text{ L} \cdot \text{atm}$ . If the pressure were increased to 1.00 atm:  $0.164 \text{ L} \cdot \text{atm} = (1.00 \text{ atm}) \times V$ ; therefore  $V = 0.164 \text{ L}$ .
- 1.12.**  $V = R \cdot \frac{nT}{p}$ , which rearranges to  $R = \frac{pV}{nT}$ .  
Therefore:  $R = \frac{(2.66 \text{ bar})(27.5 \text{ L})}{(1.887 \text{ mol})(466.9 \text{ K})} = 0.0830 \frac{\text{L} \cdot \text{bar}}{\text{mol} \cdot \text{K}}$

# Physical Chemistry David W Ball Solution Manual

**J Ma**



**Physical Chemistry David W Ball Solution Manual:**

## **Physical Chemistry David W Ball Solution Manual** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has be apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Physical Chemistry David W Ball Solution Manual**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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