

Quantum Number Problems

Use the rules for quantum numbers and notes to answer the questions below:

Rules: (1) $n = 1, 2, \dots, \infty$

(2) $l = 0, 1, 2, 3, \dots, n-1$ or s ($l = 0$), p ($l = 1$), d ($l = 2$), f ($l = 3$)

(3) $m_l = -l, -(l-1), \dots, 0, \dots, +(l-1), +l$ (e.g., when $l = 1$, $m_l = -1, 0, +1$)

(4) The number of nodal planes (N_p) is equal to the angular momentum quantum number. This means $l = N_p$.

(5) The total number of nodes is $N_T = N_p + N_r$, where N_r is the number of spherical nodes (radial nodes).

$N_T = n - 1$ where n is the principal quantum number. So $n = N_p + N_r + 1$

From these orbitals: 2s 3d 4p 5f 6d 7p

Fill in the blanks with the correct answer(s):

(answers may be used more than once or not at all; none is a possible answer)

- | | | | |
|--|-------|---|-------|
| 1. The orbital(s) with $n = 6$ | _____ | 4. The orbital(s) with 0 planar nodes | _____ |
| 2. The orbital(s) with $l = 1$ | _____ | 5. The orbital(s) with 3 spherical nodes | _____ |
| 3. The orbital(s) that can have $m_l = -2$ | _____ | 6. The orbital(s) with a total of 5 nodes | _____ |

Problems On Quantum Numbers

G Orfield



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