

Chapter 12: Atoms and the Periodic Table
Subatomic Particles

Three fundamental particles of the atom are the _____, _____, and _____. At the center of each atom lies the atomic _____, which consists of _____ and _____. The atomic number refers to the number of _____ in the nucleus. All atoms of the same element have the same number of _____, hence, the same atomic number.

Isotopes are atoms that have the same number of _____ but a different number of _____. An isotope is identified by its atomic mass number, which is the total number of _____ and _____ in the nucleus. A carbon isotope that has 6 _____ and 6 _____ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 _____ and 8 _____, on the other hand, is carbon-14.

1. Complete the following table:

| Isotope | Number of... | | |
|--------------|--------------|---------|----------|
| | Electrons | Protons | Neutrons |
| Hydrogen-1 | 1 | | |
| Chlorine-36 | | 17 | |
| Nitrogen-14 | | | 7 |
| Potassium-40 | 19 | | |
| Arsenic-75 | | 33 | |
| Gold-197 | | | 118 |

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



Physical Science Chapter 12 Answers

David Baud



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