

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \qquad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \qquad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \qquad E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left(\begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right) \qquad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \qquad PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \qquad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \qquad P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \qquad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \qquad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} \qquad M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left(\begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left(\begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right) \qquad K_w = [\text{H}^+][\text{OH}^-]$$

$$\left(\begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left(\begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right) \qquad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \qquad \text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left(\begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left(\begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right) \qquad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left(\begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left(\begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right) \qquad \Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

Sample For Staar Testing Questions For Chemistry

Guillaume Favre



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The ^AOxford Handbook of Assessment Policy and Practice in Music Education, Volume 2 Timothy Brophy,2019-01-02 In this two volume Handbook contributors from across the globe provide expert perspectives on the assessment measurement and evaluation of student learning in music

Staar Eoc Chemistry Assessment Flashcard Study System Mometrix Media,2010

STAAR Grade 8 Math for Beginners Reza Nazari,2022-01-10 A Perfect STAAR Grade 8 Math Prep Book

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SWYK on STAAR Science Gr. 8, Student Workbook Show What You Know Publishing,2013-03-01 Assess student knowledge of the Texas Essential Knowledge and Skills TEKS for Science with two full length Assessments for each subject Questions provide students with the necessary practice needed to achieve academic success on STAAR Chapters on test taking strategies and test anxiety build students confidence and test taking skills Glossaries familiarize students with vocabulary terms and concepts found on state proficiency tests Answers are provided in the Parent Teacher Edition only

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Review for the State of Texas Assessments of Academic Readiness Mometrix Media LLC, Staar Exam Secrets Test Prep, 2014-03-31 STAAR Success Strategies EOC Chemistry helps you ace the State of Texas Assessments of Academic Readiness without weeks and months of endless studying Our comprehensive STAAR Success Strategies EOC Chemistry study guide is written by our exam experts who painstakingly researched every topic and concept that you need to know to ace your test Our original research reveals specific weaknesses that you can exploit to increase your exam score more than you've ever imagined STAAR Success Strategies EOC Chemistry includes The 5 Secret Keys to STAAR Success Time is Your Greatest Enemy Guessing is Not Guesswork Practice Smarter Not Harder Prepare Don't Procrastinate Test Yourself A comprehensive General Strategy review including Make Predictions Answer the Question Benchmark Valid Information Avoid Fact Traps Milk the Question The Trap of Familiarity Eliminate Answers Tough Questions Brainstorm Read Carefully Face Value Prefixes Hedge Phrases Switchback Words New Information Time Management Contextual Clues Don't Panic Pace Yourself Answer Selection Check Your Work Beware of Directly Quoted Answers Slang Extreme Statements Answer Choice Families Along with a complete in depth study guide for your specific STAAR exam and much more **The Lancet London**

, 1868 **Current Index to Journals in Education**, 1974 **Yearbook on International Communist Affairs** Richard Felix Staar, Milorad M. Drachkovitch, Lewis H. Gann, 1971 ISI Campus Report, 1964 *The Middle East, Abstracts and Index*, 1990 The Christian Science Monitor Index, 1983 **The Wall Street Journal**, 1992 Wall Street Journal Index, 1991

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<https://crm.avenza.com/files/publication/Documents/North%20West%20Vaal%20Triangle%20Campus%20Prospectus%202016.pdf>

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