

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

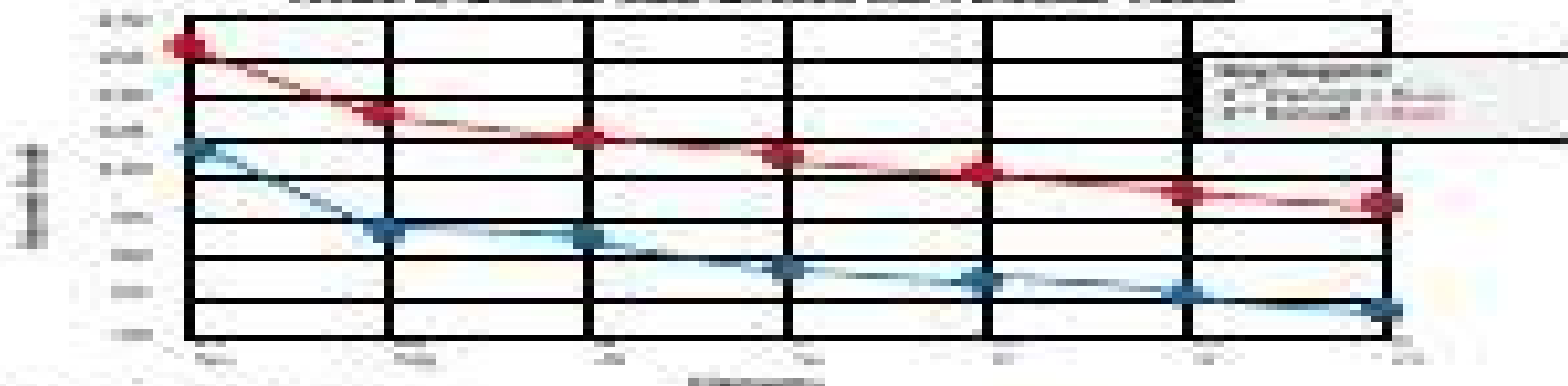
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not use the other graphs
3. Label the box as 'and include ionization
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9F	10Ne	11Na	12Mg	13Al	14Si	15P	16S
3	17Cl	18Ar	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La
8	57Ce	58Pr	59Nd	60Pm	61Sm	62Eu	63Gd	64Tb
9	65Dy	66Ho	67Er	68Tm	69Yb	70Lu	71Hf	72Ta
10	73W	74Re	75Os	76Ir	77Pt	78Au	79Hg	80Tl
11	81Pb	82Bi	83Po	84At	85Astatine	86Rn	87Fr	88Ra
12	89Ac	90Th	91Pa	92U	93Np	94Pu	95Am	96Cm
13	97Bk	98Cf	99Es	100Fm	101Md	102No	103Lr	104Ununquadium

Now label the elements above in the following way:

Trend in Atomic Size across the Periodic Table



Illustrated Trend Questions

1. Following the trend in atomic size across the periodic table decreases
2. This means two atoms that are most reactive elements (fluorine and oxygen) making themselves almost as electronegative as they can

Worked Trend Questions

1. Fluorine is the most electronegative element (F) i.e. the most electronegative element on the periodic table fluorine
2. This means fluorine that is most reactive elements (fluorine) (highest electronegativity) making them extremely electronegative

Periodicity Activity Okstate Answer Key

R Pring



Periodicity Activity Okstate Answer Key:

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