

## CHAPTER 6

## STUDY GUIDE

## The Periodic Table and Periodic Law

## Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

<del>octaves</del>	<del>atomic mass</del>	<del>atomic number</del>	<del>size</del>
<del>elements</del>	<del>properties</del>	<del>Henry Mosely</del>	<del>eight</del>
protons	periodic law	Dmitri Mendeleev	accepted

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Be	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Mosely. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

# Periodic Table Study Guide Answers

**L Towne**



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**Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011

**Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry** Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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