

Annotated periodic table of the elements

1 H1.008Hydrogen																		18 Ar39.95Argon	
3 Li6.94Lithium		4 Be9.012Beryllium																	19 K39.10Potassium
11 Na22.99Sodium		12 Mg24.31Magnesium																	21 Sc44.96Scandium
19 K39.10Potassium		20 Ca40.08Calcium																	29 Cu63.55Copper
37 Rb85.47Rubidium		38 Sr87.62Strontium																	47 Ag107.87Silver
55 Cs132.91Cesium		56 Ba137.33Barium																	65 Gd157.25Gadolinium
87 Fr223Francium		88 Ra226Radium																	103 Lr103Lanthanum

Group name*

1Hydrogen & the alkali metals

2Alkaline earth metals

11Coinage metals

12Volatile metals

13Halogens

14Crystallogens

15Pnictogens

16Chalcogens

17Halogens

18Hassium group

*Groups 1 to 10 are named after their first members i.e. Group 1 is the Scandium group

		13		14		15		16		17		18	
		5 B10.81Boron	6 C12.01Carbon	7 N14.01Nitrogen	8 O16.00Oxygen	9 F18.99Fluorine			10 Ne20.18Neon				
		13 Al26.98Aluminium	14 Si28.09Silicon	15 P30.97Phosphorus	16 S32.06Sulfur	17 Cl35.45Chlorine			18 Ar39.95Argon				
		21 Sc44.96Scandium	22 Ti47.87Titanium	23 V50.94Vanadium	24 Cr52.00Chromium	25 Mn54.94Manganese	26 Fe55.85Iron	27 Co58.93Cobalt	28 Ni58.69Nickel	29 Cu63.55Copper	30 Zn65.38Zinc	31 Ga69.72Gallium	32 Ge72.63Germanium
		39 Y88.91Yttrium	40 Zr91.22Zirconium	41 Nb92.91Niobium	42 Mo95.94Molybdenum	43 Tc98.91Technetium	44 Ru101.07Ruthenium	45 Rh101.07Rhodium	46 Pd106.32Palladium	47 Ag107.87Silver	48 Cd112.41Cadmium	49 In114.82Indium	50 Sn118.71Tin
		57 to 70 Lanthanides	71 Lu174.97Lutetium	72 Hf178.49Hafnium	73 Ta180.95Tantalum	74 W183.84Tungsten	75 Re186.21Rhenium	76 Os190.23Osmium	77 Ir192.22Iridium	78 Pt195.08Platinum	79 Au196.97Gold	80 Hg200.59Mercury	81 Tl204.38Thallium
		89 to 102 Actinides	103 Lr103Lanthanum	104 Rf181Rutherfordium	105 Db185Dubnium	106 Sg271Seaborgium	107 Bh272Bohrium	108 Hs277Hassium	109 Mt276Meitnerium	110 Ds281Darmstadtium	111 Rg282Roentgenium	112 Cn285Copernicium	113 Nh286Nihonium
			103 Lr103Lanthanum	104 Rf181Rutherfordium	105 Db185Dubnium	106 Sg271Seaborgium	107 Bh272Bohrium	108 Hs277Hassium	109 Mt276Meitnerium	110 Ds281Darmstadtium	111 Rg282Roentgenium	112 Cn285Copernicium	113 Nh286Nihonium
			114 Fl289Flerovium	115 Mc289Moscovium	116 Lv293Livermorium	117 Ts294Tennessine	118 Og294Oganesson						

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Halogens
- Chalcogens
- Pnictogens
- Chalcogens
- Halogens
- Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
Mendeleev

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

* as liquid P, the most stable form in ambient conditions.

Group 1

This group is Sc-Y, and the lanthanides (La-Lu) and the actinides (Ac-Lu). Trends going from Sc-Y and the lanthanides (La-Lu) trends going from Sc-Y and the actinides (Ac-Lu) trends going from Sc-Y and the actinides (Ac-Lu).

Nonmetals

Chalcogens
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal.

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Michael Brown



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