

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

Name _____

Date _____

- 2 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 4 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 7.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 58.8 grams. Calculate the percent yield.

- 6 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). Calculate the theoretical yield.



What is the theoretical yield?

- 7 1.8 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 10.5 grams. Calculate the percent yield of phosphorus trichloride.

Name _____

Date _____

- 8 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 23.02 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCl}_2}{100.09 \text{ g CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 23.02 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 9 The actual yield was 22.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{22.5 \text{ g}}{23.02 \text{ g}} \times 100 = 97.7\%$$

- 10 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 5.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 5.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 15.0 \text{ g KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

YIELD

ANSWER KEY INCLUDED

Percent Yield Example Answer Key

**National Center for Health Statistics
(U.S.)**



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1. Understanding the eBook Percent Yield Example Answer Key
 - The Rise of Digital Reading Percent Yield Example Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Percent Yield Example Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Percent Yield Example Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Yield Example Answer Key
 - Personalized Recommendations
 - Percent Yield Example Answer Key User Reviews and Ratings
 - Percent Yield Example Answer Key and Bestseller Lists
5. Accessing Percent Yield Example Answer Key Free and Paid eBooks
 - Percent Yield Example Answer Key Public Domain eBooks
 - Percent Yield Example Answer Key eBook Subscription Services

- Percent Yield Example Answer Key Budget-Friendly Options
- 6. Navigating Percent Yield Example Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Percent Yield Example Answer Key Compatibility with Devices
 - Percent Yield Example Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percent Yield Example Answer Key
 - Highlighting and Note-Taking Percent Yield Example Answer Key
 - Interactive Elements Percent Yield Example Answer Key
- 8. Staying Engaged with Percent Yield Example Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Percent Yield Example Answer Key
- 9. Balancing eBooks and Physical Books Percent Yield Example Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percent Yield Example Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Percent Yield Example Answer Key
 - Setting Reading Goals Percent Yield Example Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Percent Yield Example Answer Key
 - Fact-Checking eBook Content of Percent Yield Example Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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