

**Assessment 17 Worksheet: Redox Reactions & Oxidation Numbers**

Complete the following questions, and upload the file to the dropbox.

1. Why can an oxidation reaction not happen without a reduction reaction? Your response should clearly communicate your understanding of the nature of redox reactions. (1 point)  
*An oxidation reaction can not happen without a reduction reaction because for an oxidation reaction there must also be a reduction happening at the same time. This is because oxidation happens because of the "compound" that causes the oxidation, this "compound" must be reduced otherwise oxidation will not occur.*

2. Determine the oxidation number of the specified element in each of the following. (6 points)

- a. Nitrogen in  $\text{NF}_3$                       +3
- b. Sulfur in  $\text{S}_8$                               0
- c. Chromium in  $\text{CrO}_4^{2-}$                       +5
- d. Phosphorus in  $\text{P}_2\text{O}_5$                       +5
- e. Carbon in  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$                       0
- f. Carbon in  $\text{CHCl}_3$                           +2

3. Determine the oxidation number for each element in the following substances. (6 points)

- a.  $\text{H}_2\text{SO}_3$               hydrogen: +1                      sulfur: +4                      oxygen: -2
- b.  $\text{OH}^-$                       oxygen: -2                      hydrogen: +1
- c.  $\text{HPO}_4^{2-}$               hydrogen: +1                      phosphorus: +5                      oxygen: -2

4. Determine whether each reaction is a redox reaction by finding the oxidation numbers. Show the oxidation numbers to justify your answer. (6 points)



# Redox Reactions Chapter Assessment Answers

**Katrin Zwirglmaier**



## **Redox Reactions Chapter Assessment Answers:**

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