

Practice Problems: Redox Reactions (Answer Key)

1. Determine the oxidation number of the elements in each of the following compounds:



H: +1, O: -2, C: +4



N: 0



Zn: 2+, H: +1, O: -2



N: +3, O: -2



Li: +1, H: -1



Fe: +8/3, O: -2

2. Identify the species being oxidized and reduced in each of the following reactions:



Cr^+ : oxidized, Sn^{4+} : reduced



Hg^{2+} : reduced, Fe: oxidized



As: oxidized, Cl_2 : reduced

3. Would you use an oxidizing agent or reducing agent in order for the following reactions to occur?

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