# REDOX REACTIONS

Name \_\_\_\_

Oxidizing agent

For the equations below, identify the substance oxidized, the substance reduced, the oxidizing agent, the reducing agent, and write the oxidation and reduction half reactions.

1. 2H, + 0, - 2H,0 OK: 2Hb > 4H + 4e Loducing agent RED: 0, +4e -> 20^2 - Oxidizing agent 2. Fe + Zn - Fe + Zn SX: Fe -> Fe 2+ + 2e Leducing agent

3.  $2AI + 3Fe^{-1} - 2AI^{-1} + 3Fe$   $OX : \partial AI^{0} - > \partial AI^{13} + 6e^{-}$  laducing agent  $Red : 3Fe^{+2} + 6e^{-} - > 3Fe$  Oxidizing agent

4. Cu + 2AgNO, + CUKNO), + 2Ag

OX: Cu° -> Cu+2 + 2e reducing agent

red: 2Ag+ + 2e -> 2Ag° oxidizing agent

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Mr. Ajeet Singh

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