

# Redox Exercise (Part 2)

## Chapter 20 Worksheet: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.



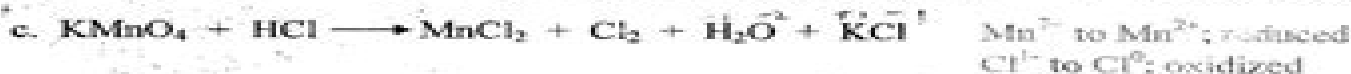
11. Give the oxidation number of each kind of atom or ion.

a. sulfate	b. Sn	c. $\text{S}^{2-}$	d. $\text{Fe}^{3+}$	e. $\text{Sn}^{4+}$	f. nitrate	g. ammonium
2-	0	2-	3+	4+	1-	1+

12. Calculate the oxidation number of chromium in each of the following.

a. $\text{Cr}_2\text{O}_3$	b. $\text{Na}_2\text{Cr}_2\text{O}_7$	c. $\text{CrSO}_4$	d. chromate	e. dichromate
3+	6+	2+	6+	6+

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)



14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- |                               |                          |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur    | reducing agent: carbon   |
| b. oxidizing agent: nitrogen  | reducing agent: iodine   |
| c. oxidizing agent: manganese | reducing agent: chloride |
| d. oxidizing agent: nitrogen  | reducing agent: antimony |

# Redox Equations With Answers Key

**PT Brinkman**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the PT Brinkman text.

## **Redox Equations With Answers Key:**

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