

## Chapter 6 Chemical Bonds

**Section 6.2 Covalent Bonding****(pages 165–169)**

*This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.*

**Reading Strategy (page 165)**

**Relating Text and Visuals** As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

| Molecular Models   |             |
|--------------------|-------------|
| Model              | Description |
| Electron dot       |             |
| Structural formula |             |
| Space-filling      |             |
| Electron cloud     |             |

**Covalent Bonds (pages 165–167)**

- Describe a covalent bond. \_\_\_\_\_  
\_\_\_\_\_
- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
  - electron dot
  - structural formula
  - space-filling
  - electron cloud
- Describe a molecule. \_\_\_\_\_  
\_\_\_\_\_
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. \_\_\_\_\_
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
  - zero
  - one
  - two
  - three

# Note Taking Guide Covalent Bonding Answer Key

**D Kirk**



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