

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H_2O , which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less H_2O (12.0 g) in our initial amount so
 H_2O is the L.R

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g $\text{Mg}(\text{OH})_2$. What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

Sample Problems H Gas Stoichiometry Answers

JS Bruner



Sample Problems H Gas Stoichiometry Answers:

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Sample Problems H Gas Stoichiometry Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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