#### Quantum Numbers

- Used to an in an
  - 1
  - Represents \_\_\_\_\_\_ energy level of \_\_\_\_\_\_\_

\_\_\_\_\_ ar of \_\_\_\_\_ in an

Example: What is the maximum number of electrons that can be in the \_\_\_\_\_ main energy level?

£

- The \_\_\_\_\_\_
- Describes the \_\_\_\_\_\_ within an \_\_\_\_\_\_
- of arbital \_\_\_\_\_possible is \_\_\_\_\_\_

#### Orbital Shopes

designated\_\_\_\_\_

- level 21
- level 3
- level 4:

#### How many electrons can each sublevel hold?

- s = 1 orbital x 2 e /orbital = \_\_\_\_e
- p = 3 orbitals x 2 e /orbital = \_\_\_\_e
- d = 5 orbitals x 2 e /orbital = \_\_\_\_e
  - f = 7 orbitals x 2 e /orbital = \_\_\_\_e

# **Note Taking Guide Episode 304 Key**

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