

Chemical Formula

Telling this problem involves and relates all atoms present in a compound.

Molecular $\text{C}_{10}\text{H}_{18}\text{O}$ (molecular)

Formula represents the number of atoms of each element in a single molecule.

Empirical $\text{C}_5\text{H}_9\text{O}$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound.

Atomic Mass

The average mass of all atoms of an element is 12.01 amu.

Mass

$1\text{ kg} = 1000\text{ g}$ (mass)

$1\text{ mole} = 6.022 \times 10^{23}$ (mole)

Substance

$1\text{ mole} = 6.022 \times 10^{23}$ (mole)

$1\text{ mole} = 6.022 \times 10^{23}$ (mole)

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Yicheng Fang



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