

No Worst, There Is None

No worst, there is none. Pitched past pitch of grief,
More pangs will, schooled at forepangs, wilder wring.
Comforter, where, where is your comforting?
Mary, mother of us, where is your relief?
My cries heave, herds-long; huddle in a main, a chief{\-}
Woe, w{'o}rld-sorrow; on an {'a}ge-old {'a}nvil w{'i}nce and s{'i}ng --

Then lull, then leave off. Fury had shrieked "No ling-
Ering! Let me be fell: force I must be brief."

O the mind, mind has mountains; cliffs of fall
Frightful, sheer, no-man-fathomed. Hold them cheap
May who ne'er hung there. Nor does long our small
Durance deal with that steep or deep. Here! creep,
Wretch, under a comfort serves in a whirlwind: all
Life death does end and each day dies with sleep.

Gerard Manley Hopkins

No Worst There Is None

Minjie Lin



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of h_2so_4 and 1 mole of oh^- combine to form 2 moles of h_2o $\Delta h = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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