- What is the most stable group on the Periodic Table? Why is this group stable?
 Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full.
- Name the most reactive group(s) on the Periodic Table? Why are they reactive?
 Alkali Metals (Group 1) & Halogens (Group 17); Alkali Metals want to lose one valence electron in order to become stable and Halogens want to gain one valence electron to be stable.
- 3. In the boxes below draw the Electron dot diagrams for:

Sodium	Calcium	Boron	Chlorine	Krypton
Na-	Ca-	-B-	: CI-	: Kr :
	10 Ca			

4. Name the three (3) types of chemical bonds:

Ionic bonds Covalent bonds Polyatomic

5. Define ionic bond:

The force that holds cations and anions together.

6. Define covalent bond:

A chemical bond in which two atoms share a pair of valence electrons.

7. Define polyatomic ion:

A covalently bonded group of atoms that has a positive or negative charge and acts as a unit.

- 8. In an ionic bond, electrons are transferred,
- 9. Who are the participants in an ionic bond? Metals and nonmetals
- 10. An ion with a positive charge is a cation; on ion with a negative charge is an anion.
- 11. In a covalent bond, electrons are shared.
- 12. Who are the participants in a covalent bond? Two nonmetals
- Identify whether the following compounds are covalent (two nonmetals), ionic (a nonmetal and a metal), or polyatomic (both).

(O	CoCla	IONIC (transferred)
ъ.	H ₂ O	COVALENT (shared)
C.	BaSO ₄	POLYATOMIC both)
d.	CH4	COVALENT (shared)
e	LiBr	IONIC (transferred)
f.	NH4CI	POLYATOMIC (both)
g.	KI	IONIC (transferred)

Review Chapter 6 Chemical Bonds Physical Science

Robert R. Redfield

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