

Answers to the Gases Review Sheet

1. 3.21 mm Hg = 0.321 cm Hg = 0.00422 atm
2. 6.9 kPa = 6900 Pa = 52 mm Hg
3. 436 mm Hg = 0.574 atm = 58.1 kPa
4. 2.9 kPa = 22 torr = 22 mm Hg
5. A sample of air has a volume of 2.25 cm³ when its temperature is 298 K. If the temperature is increased to 373 K without changing the pressure, what is the new volume of the sample in air?

$$\begin{aligned} V_1 &= 2.25 \text{ cm}^3 & \frac{V_1}{T_1} &= \frac{V_2}{T_2} & V_2 &= \frac{V_1 T_2}{T_1} \\ T_1 &= 298 \text{ K} \\ T_2 &= 373 \text{ K} \\ V_2 &= x \end{aligned}$$

Charles's

$$V_2 = \frac{2.25 \text{ cm}^3 (373 \text{ K})}{298 \text{ K}} = \mathbf{2.82 \text{ cm}^3}$$

6. As the temperature of a sample of nitrogen increases from 273 K, the volume of the sample changes from 275 mL to 325 mL. Assuming the pressure is constant, what is the final temperature of the gas?

$$\begin{aligned} V_1 &= 275 \text{ mL} & \frac{V_1}{T_1} &= \frac{V_2}{T_2} & T_2 &= \frac{T_1 V_2}{V_1} \\ T_1 &= 273 \text{ K} \\ T_2 &= x \\ V_2 &= 325 \text{ mL} \end{aligned}$$

Charles's

$$T_2 = \frac{325 \text{ mL} (273 \text{ K})}{275 \text{ mL}} = \mathbf{323 \text{ K}}$$

7. The pressure of a 5.71 L sample of neon gas is 23.4 kPa. Calculate the new pressure when the volume becomes 3.40 L.

$$\begin{aligned} V_1 &= 5.71 \text{ L} & P_1 V_1 &= P_2 V_2 & P_2 &= \frac{P_1 V_1}{V_2} \\ P_1 &= 23.4 \text{ kPa} \\ P_2 &= x \end{aligned}$$

$$V_2 = 3.40 \text{ L} \quad P_2 = \frac{5.71 \text{ L} (23.4 \text{ kPa})}{3.40 \text{ L}} = 39.3 \text{ kPa} = \mathbf{0.388 \text{ atm}}$$

Boyle's

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