

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 11 REVIEW***Gases***SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ The molar mass of a gas at STP is the density of that gas  
(a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L.  
(b) divided by the mass of 1 mol. (d) divided by 22.4 L.
2. \_\_\_\_\_ For the expression  $V = \frac{nRT}{P}$ , which of the following will cause the volume to increase?  
(a) increasing  $P$  (c) increasing  $T$   
(b) decreasing  $T$  (d) decreasing  $n$
3. Two sealed flasks, A and B, contain two different gases of equal volume at the same temperature and pressure.  
\_\_\_\_\_ a. The two flasks must contain an equal number of molecules. True or False?  
\_\_\_\_\_ b. The two samples must have equal masses. True or False?

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

4. Use the data in the table below to answer the following questions.

Formula	Molar mass (g/mol)
$N_2$	28.02
CO	28.01
$C_2H_2$	26.04
He	4.00
Ar	39.95

(Assume all gases are at STP.)

- \_\_\_\_\_ a. Which gas contains the most molecules in a 5.0 L sample?  
\_\_\_\_\_ b. Which gas is the least dense?  
\_\_\_\_\_ c. Which two gases have virtually the same density?  
\_\_\_\_\_ d. What is the density of  $N_2$  measured at STP?

# Review Gases Section 3 Short Answer Key

**RS Peters**



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