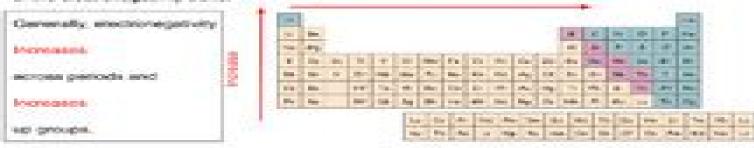
Periodicity - Electronegativity

Number of Co.

Angeler Key

 Fill in the blanks below, and then dose arrows on to the periodic table indicating the direction of the electromagnitudy trend.



List the electionts before in order of increasing electronegativity. (Smallest to largest).

| Ba, Ra, Ca, Be, Mg | Zn. Nr. Cu. Sc. Ti | F. B. O. Rb. Fr |
|-----------------------|----------------------|--------------------|
| Ro, Ba, Ca, Mg, Be | Se, S. Nr. Cu. Zn | Fr. Rb. B. O. F |
| No., Sn., Rb., Fr., I | F, S.A. Ele, M., Pda | P. St. Ar, Mg. Cu. |
| Fr. Rb., No., Sn., I | K., Nie, S.I. Ber, F | Cu. Mg. St. P. Ar |

Are the elements before in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

| Ca. K. Se. Sr. Sr. Se, Ca. K. | E. St. CL F | O. B. Li, No. |
|----------------------------------|-----------------|--------------------------------|
| Y. Se, Rb, Fr. | All, Sc. St. Bo | Au. Pb. To. S S. Te. Pb. Au |

4. Briefly explain why electronogativity changes up groups.

Generally, atactronogistivity increases as you move up in a group. This is because the elemic number decreases, and as a result, there is less discarce between the nucleus and valence excitents. In other words, there is a smaller atomic radius.

5. Briefly explain why electronegativity changes across periods.

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased altraction to electrons.

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