

# Graphing Periodic Trends Activity

Name: \_\_\_\_\_ Period: \_\_\_\_\_

Use the information in these tables to complete the graph as described below

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| H      | 31                         | 1312                                      | 2.1                               |
| Li     | 128                        | 520                                       | 1.0                               |
| Na     | 166                        | 496                                       | 0.9                               |
| K      | 203                        | 410                                       | 0.8                               |
| Rb     | 220                        | 403                                       | 0.8                               |
| Cs     | 244                        | 376                                       | 0.7                               |

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| Na     | 166                        | 496                                       | 0.9                               |
| Mg     | 141                        | 738                                       | 1.2                               |
| Al     | 121                        | 578                                       | 1.5                               |
| Si     | 111                        | 787                                       | 1.8                               |
| P      | 107                        | 1012                                      | 2.1                               |
| S      | 105                        | 1000                                      | 2.5                               |
| Cl     | 102                        | 1251                                      | 3.0                               |
| Ar     | 106                        | 1521                                      | ---                               |

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

|                         |                                                                                                                                                                                          |    |    |   |   |    |    |
|-------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----|----|---|---|----|----|
| H                       | <p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p> |    |    |   |   |    |    |
| Li                      |                                                                                                                                                                                          |    |    |   |   |    |    |
| Na<br>166<br>496<br>0.9 | Mg                                                                                                                                                                                       | Al | Si | P | S | Cl | Ar |
| K                       | <p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>          |    |    |   |   |    |    |
| Rb                      |                                                                                                                                                                                          |    |    |   |   |    |    |
| Cs                      |                                                                                                                                                                                          |    |    |   |   |    |    |

# Periodic Trends Graphing Activity Answer Key

**D Keegan**



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## **Table of Contents Periodic Trends Graphing Activity Answer Key**

1. Understanding the eBook Periodic Trends Graphing Activity Answer Key
  - The Rise of Digital Reading Periodic Trends Graphing Activity Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Graphing Activity Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Graphing Activity Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Graphing Activity Answer Key
  - Personalized Recommendations
  - Periodic Trends Graphing Activity Answer Key User Reviews and Ratings
  - Periodic Trends Graphing Activity Answer Key and Bestseller Lists
5. Accessing Periodic Trends Graphing Activity Answer Key Free and Paid eBooks
  - Periodic Trends Graphing Activity Answer Key Public Domain eBooks
  - Periodic Trends Graphing Activity Answer Key eBook Subscription Services
  - Periodic Trends Graphing Activity Answer Key Budget-Friendly Options

6. Navigating Periodic Trends Graphing Activity Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Graphing Activity Answer Key Compatibility with Devices
  - Periodic Trends Graphing Activity Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Graphing Activity Answer Key
  - Highlighting and Note-Taking Periodic Trends Graphing Activity Answer Key
  - Interactive Elements Periodic Trends Graphing Activity Answer Key
8. Staying Engaged with Periodic Trends Graphing Activity Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Graphing Activity Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Graphing Activity Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Graphing Activity Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Graphing Activity Answer Key
  - Setting Reading Goals Periodic Trends Graphing Activity Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Graphing Activity Answer Key
  - Fact-Checking eBook Content of Periodic Trends Graphing Activity Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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