



Percent Composition Of Compounds

DATE _____

Find the percent composition in the following compounds.

1. H₂O

$$\text{molar mass: } 2 + 1.01 + 16.00 = 18.02 \text{ g mol}^{-1}$$

$$\%H = \frac{2 + 1.01}{18.02} = 100\% = 11.2\%$$

$$\%O = \frac{16.00}{18.02} = 100\% = 88.8\%$$

2. CO₂

$$\text{molar mass: } 12.01 + 2 + 16.00 = 44.01 \text{ g mol}^{-1}$$

$$\%C = \frac{12.01}{44.01} = 100\% = 27.3\%$$

$$\%O = \frac{2 + 16.00}{44.01} = 100\% = 72.7\%$$

3. NaOH

$$\text{molar mass: } 22.99 + 16.00 + 1.01 = 40.00 \text{ g mol}^{-1}$$

$$\%Na = \frac{22.99}{40.00} = 100\% = 57.5\%$$

$$\%H = \frac{1.01}{40.00} = 100\% = 2.5\%$$

$$\%O = \frac{16.00}{40.00} = 100\% = 40.0\%$$

4. H₂SO₄

$$\text{molar mass: } 2 + 1.01 + 32.07 + 4 + 16.00 = 98.07 \text{ g mol}^{-1}$$

$$\%H = \frac{2 + 1.01}{98.07} = 100\% = 2.1\%$$

$$\%S = \frac{32.07}{98.07} = 100\% = 32.7\%$$

$$\%O = \frac{4 + 16.00}{98.07} = 100\% = 65.2\%$$

5. Ca(OH)₂

$$\text{molar mass: } 40.08 + 2 + 16.00 + 2 + 1.01 = 74.10 \text{ g mol}^{-1}$$

$$\%Ca = \frac{40.08}{74.10} = 100\% = 54.1\%$$

$$\%H = \frac{2 + 1.01}{74.10} = 100\% = 2.7\%$$

$$\%O = \frac{2 + 16.00}{74.10} = 100\% = 22.3\%$$

Percentage Composition Chemistry Answerkey

J Ma



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Percentage Composition Chemistry Answerkey Book Review: Unveiling the Power of Words

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