



100

1	2	3	4
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PHYSICAL SCIENCES 23
(CHEMISTRY)
EXAMPLE

Abstract



The quarterly sales volume of Pappas Printing is as follows:

QUESTION 3

**PHYSICAL SCIENCES
PAPER 2**

INTERMOLECULAR FORCES

GRADE 11**NOVEMBER 2020**

Figure 1

[illegible]

- 12.1 The leading parts of the table below were obtained by applying the asymptotic normality to compute the leading parts of group 1's density. The columns will provide you information for the other subgroups. **1** and **2** correspond to

Treatment		Mineral loss (g/100 g)	Salinity (%)
1	Control	10.5	0.1
2	100	9.1	0.1

- | | | |
|------|--|----|
| Q1.1 | Define the term <i>solubility</i> . | 10 |
| Q1.2 | Is each phase in equilibrium (A and B) or (C)? Justify why or explain your work! | 10 |
| Q1.3 | Draw the type of intermolecular force that exists between molecules of each component A and B for A-B pairs within A and B mixtures. | 10 |
| Q1.4 | What kind of the components (A or B) has H-bonding capacity? Draw hydrogen bond! | 10 |
| Q1.5 | How is water to get frozen? | 10 |
| Q1.6 | Which component (A or B) will have the higher vapor pressure? | 10 |
| Q1.7 | How is benzene to be oxidized to carbon dioxide by using the table? | 10 |

Structure 1 (95% bootstrap in the cells) has a smaller minimum value than both competitors 2 and 3 but has a relatively higher average value (Table 2).

- 2014** Explain why the ending point of compound II is HCO_2H instead of compound II and III by referring to the FMOs and electronegativity of the substituents. (10 points)

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- Q102: Explain the concept of "SMART" goals and its importance in project management.

100

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Physical Science Paper 2 Intermolecular Forces Investigation

Patrick Vollmar



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