

ATOMIC STRUCTURE AND THE PERIODIC TABLE CHAPTER 4 WORKSHEET

PART A

Given the following isotopes, determine the atomic number, the mass number, the number of protons, electrons and neutrons.

Isotope Symbol	Atomic Number	Mass Number	Protons	Electrons	Neutrons	Isotope Name
$^{131}_{53}\text{I}$	53	131	53	53	78	iodine - 131
$^{35}_{16}\text{S}$	16	35	16	16	19	sulphur - 35
^4_2He	2	4	2	2	2	helium - 4
$^{27}_{13}\text{Al}$	13	27	13	13	14	aluminium - 27
$^{81}_{36}\text{Kr}$	36	81	36	36	45	krypton - 81
$^{81}_{37}\text{Rb}$	37	81	37	37	44	rubidium - 81

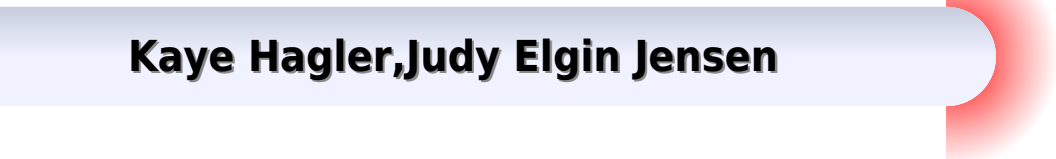
PART B

Complete the following chart by writing the symbol for the isotope of the following elements. In addition, give the number of protons, electrons, mass number and atomic number and complete the element name.

Element Name	Neutrons	Protons	Electrons	Mass Number	Atomic Number	Isotope Symbol
Uranium- 237	145	92	92	237	92	$^{237}_{92}\text{U}$
Chlorine- 45	28	17	17	45	17	$^{45}_{17}\text{Cl}$
Oxygen- 17	9	8	8	17	8	$^{17}_8\text{O}$
Boron- 11	6	5	5	11	5	$^{11}_5\text{B}$
Beryllium- 9	5	4	4	9	4	^9_4Be
Hydrogen- 2	1	1	1	2	1	^2_1H
Carbon- 14	8	6	6	14	6	$^{14}_6\text{C}$

Physical Science Chapter 4 Answers

Kaye Hagler, Judy Elgin Jensen



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