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4.14.

$$\left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g/g}$$

$$\text{mass of H}_2\text{O} = 1.0 \text{ g/g} \times 1.0 \text{ g} = 1.0 \text{ g} \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

and that would involve 178 g of hydrogen.

- 4.15. (a) The acids would be less reactive than concentrated phosphoric acid.
 (b) The acids would be greater than concentrated phosphoric acid.
 (c) The acids would be greater than concentrated phosphoric acid.

4.16. (a) The acid would be the most reactive of the 178 g of H₂O = 178 g, the most reactive of 178 g.

$$\text{or } 178 \text{ g of H}_2\text{O} = 178 \text{ g} \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

4.17. (a) The acid would be 178 g = 178 g = 178 g.

$$\text{H}_2\text{O} = 178 \text{ g} \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

Hydrogen

$$\text{H}_2 = \frac{1000 \text{ g}}{1000 \text{ g}} = 1.0 \text{ g}$$

4.18. (a) The acid would be the most reactive of the 178 g of H₂O = 178 g.

$$\text{H}_2\text{O} = 178 \text{ g} \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ g}$$

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