

All quantities on the right are given to us except  $n$ , which can be computed from the given mass of Ar.

$$n = \frac{25 \text{ g}}{39.95 \text{ g mol}^{-1}} = 0.626 \text{ mol}$$

$$\text{so } p = \frac{(0.626 \text{ mol}) \times (8.31 \times 10^{-2} \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}) \times (30 + 273 \text{ K})}{1.5 \text{ dm}^3} = \boxed{10.3 \text{ bar}}$$

not 2.0 bar.

(b) The van der Waals equation [1.21a] is

$$p = \frac{RT}{V_m - b} - \frac{a}{V_m^2}$$

$$\begin{aligned} \text{so } p &= \frac{(8.31 \times 10^{-2} \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}) \times (30 + 273) \text{ K}}{(1.5 \text{ dm}^3 / 0.626 \text{ mol}) - 3.20 \times 10^{-2} \text{ dm}^3 \text{ mol}^{-1}} \\ &\quad - \frac{(1.337 \text{ dm}^6 \text{ atm mol}^{-2}) \times (1.013 \text{ bar atm}^{-1})}{(1.5 \text{ dm}^3 / 0.626 \text{ mol})^2} = \boxed{10.4 \text{ bar}} \end{aligned}$$

**E1.2(b)** (a) Boyle's law [1.5] applies.

$$pV = \text{constant} \quad \text{so} \quad p_f V_f = p_i V_i$$

This equation can be solved for either initial or final pressure, hence

$$p_i = \frac{p_f V_f}{V_i} = \frac{(1.97 \text{ bar}) \times (2.14 \text{ dm}^3)}{(2.14 + 1.80) \text{ dm}^3} = \boxed{1.07 \text{ bar}}$$

(b) The original pressure in Torr is

$$p_i = (1.07 \text{ bar}) \times \left( \frac{1 \text{ atm}}{1.013 \text{ bar}} \right) \times \left( \frac{760 \text{ Torr}}{1 \text{ atm}} \right) = \boxed{803 \text{ Torr}}$$

**E1.3(b)** The relationship between pressure and temperature at constant volume can be derived from the perfect gas law,  $pV = nRT$  [1.8]

$$\text{so } p \propto T \quad \text{and} \quad \frac{p_i}{T_i} = \frac{p_f}{T_f}$$

The final pressure, then, ought to be

$$p_f = \frac{p_i T_f}{T_i} = \frac{(125 \text{ kPa}) \times (11 + 273) \text{ K}}{(23 + 273) \text{ K}} = \boxed{120 \text{ kPa}}$$

**E1.4(b)** According to the perfect gas law [1.8], one can compute the amount of gas from pressure, temperature, and volume.

$$pV = nRT$$

$$\text{so } n = \frac{pV}{RT} = \frac{(1.00 \text{ atm}) \times (1.013 \times 10^5 \text{ Pa atm}^{-1}) \times (4.00 \times 10^3 \text{ m}^3)}{(8.3145 \text{ J K}^{-1} \text{ mol}^{-1}) \times (20 + 273) \text{ K}} = 1.66 \times 10^5 \text{ mol}$$

# Physical Chemistry 9th Solution Manual

**R Barnett**



## **Physical Chemistry 9th Solution Manual:**

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Waldeck, Horst-Dieter Försterling, 2024-10-25 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text

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