

Name: _____

CLASSIFYING CHEMICAL REACTIONS

Classify the following chemical reactions as one of the following: synthesis, decomposition, single-replacement, double-replacement, or combustion.

$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
$2\text{K} + 2\text{H}_2\text{O} \rightarrow \text{H}_2 + 2\text{KOH}$	$\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
$2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$	$\text{Ca} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2$
$2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$	$\text{Pb(NO}_3)_2 + 2\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{HNO}_3$
$\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$	$2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
$2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$	$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
$\text{Pb(NO}_3)_2 + 2\text{K} \rightarrow \text{Pb} + 2\text{KNO}_3$	$\text{H}_2\text{CO}_3 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
$2\text{Fe} + \text{O}_2 \rightarrow 2\text{FeO}$	$\text{Zn} + 2\text{HCl} \rightarrow \text{H}_2 + \text{ZnCl}_2$
$\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$

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L Reisser



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Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh, 1921 *Chemical Reactions*, 2006

THEORY CHEM REACT DYN Michael Baer, 1985-11-19 **Encyclopedia of Chemical Reactions** Carl Alfred

Jacobson, Clifford A. Hampel, 1956 **The Reaction Path in Chemistry: Current Approaches and Perspectives** D.

Heidrich, 2013-03-09 The so called reaction path RP with respect to the potential energy or the Gibbs energy free enthalpy is one of the most fundamental concepts in chemistry It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics It is particularly useful for the interpretation of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept Chemical Reactions Kaufman George Falk, 1920 **Gas-Phase**

Reactions V.N. Kondratiev, E.E. Nikitin, 2012-12-06 The present monograph appears after the death of Professor V N Kondratiev one of those scientists who have greatly contributed to the foundation of contemporary gas kinetics The most fundamental idea of chemical kinetics put forward at the beginning of the twentieth century and connected with names such as W Nernst M Bodenstein N N Semenov and C N Hinshelwood was that the complex chemical reactions are in fact a manifestation of a set of simpler elementary reactions involving but a small number of species V N Kondratiev was one of the first to adopt this idea and to start investigations on the elementary chemical reactions proper These investigations revealed explicitly that every elementary reaction in turn consisted of many elementary events usually referred to as elementary processes It took some time to realize that an elementary reaction represented in a very simple way by a macroscopic kinetic equation can be described on a microscopic level by a generalized Boltzmann equation Nevertheless up to the middle of the twentieth century gas kinetics was mainly concerned with the interpretation of complex chemical reactions via a set of elementary reactions But later on the situation changed drastically First the conditions for reducing microscopic equations to macroscopic ones were clearly set up These are essentially based on the fact that the small perturbations of the Maxwell Boltzmann distribution are caused by the reaction proper *THEORY CHEM REACT DYN* Michael Baer, 1985-11-19

Chemical Reactivity in Liquids Michael Moreau, Pierre Turq, 2012-12-06 Understanding chemical reactivity has been the permanent concern of chemists from time immemorial. If we were able to understand it and express it quantitatively, there would practically remain no unsolved mystery and reactions would be fully predictable with their products and rates and even side reactions. The beautiful developments of thermodynamics through the 19th century supplied us with the knowledge of the way a reaction progresses and the statistical view initiated by Gibbs has progressively led to an understanding closer to the microscopic phenomena. But it was always evident to all that these advances still left our understanding of chemical reactivity far behind our empirical knowledge of the chemical reaction in its practically infinite variety. The advances of recent years in quantum chemistry and statistical mechanics, enhanced by the present availability of powerful and fast computers, are very fast changing this picture and bringing us really close to a microscopic understanding of chemical equilibria, reaction rates, etc. This is the reason why our Society encouraged a few years ago the initiative of Professor Savo Bratos, who with a group of French colleagues prepared an impressive study on *Reactivite chimique en phase liquide*, a prospective report which was jointly published by the Societe Francaise de Chimie. **Chemical Triggering** Gebran J. Sabongi, 2012-12-06 Chemical reactions which can on demand be switched on and off are valuable for industrial applications. In order to make the best use of these reactions, it is essential to have them readily available for a research chemist. The chemical literature in general has not yet identified or grouped such reactions. However, their existence is relatively abundant. This book is meant as a survey of those reactions which have potential utility in industrially useful processes. These reactions are grouped under the title of *chemical release reactions*, which can be triggered by heat, light, electric current, etc. to release a specific compound from or change in the physical or chemical properties of a unimolecular reactant. The book is divided into chapters covering ways to trigger the release of certain chemicals. Each chapter is further divided into sections, each beginning with a brief introduction, of analogies of the discussed reactions and of how they were used in reported industrial processes. This survey is not meant to be absolute or exhaustive but rather to be directive, to be as complete as possible, and to provide food for further thought.

Real World Examples of Four Types of Chemical Reactions, James W. Smith developed Real World Examples of Four Types of Chemical Reactions, a physical science and chemistry lesson for 9th grade students. The students use the Internet to locate information about industrial, pharmaceutical, and environmental operations that use examples of the four general classes of chemical reactions: synthesis, decomposition, single replacement reactions, and double replacement reactions. The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K-12 classroom. Advances in Chemical Reaction Dynamics Peter M. Rentzepis, Christos Capellos, 1986-08-31 Proceedings of the NATO Advanced Study Institute, Iraklion, Crete, Greece, August 25-September 7, 1985.

Chemical Reactions in Solvents and Melts G. Charlot, B. Trémillon, 2013-10-22 Chemical Reactions in Solvents and Melts discusses the use of organic and inorganic compounds as well as of melts as solvents. This book examines the

applications in organic and inorganic chemistry as well as in electrochemistry Organized into two parts encompassing 15 chapters this book begins with an overview of the general properties and the different types of reactions including acid base reactions complex formation reactions and oxidation reduction reactions This text then describes the properties of inert and active solvents Other chapters consider the proton transfer reactions in polar solvents as well as the transfer of other ions This book discusses as well the solubility in a number of solvents by the formation of different bonds between the solute and the solvent molecule The final chapter deals with the general characteristics of the oxidation reduction reactions of melts This book is a valuable resource for chemists students and researchers

Theory of Chemical Reaction Dynamics
Antonio Laganà, 2005

Turbulent Mixing and Chemical Reactions Jerzy Bałdyga, John R. Bourne, 1999-03-12 Turbulent Mixing and Chemical Reactions Jerzy Bałdyga Warsaw University of Technology Poland John R Bourne Visiting Professor University of Birmingham UK and Emeritus Professor ETH Zurich Switzerland The way in which reagents are mixed can greatly influence the yield and range of products formed by fast multiple chemical reactions Understanding this phenomenon enables chemists to carry out reactions more selectively make better use of raw materials and simplify product workup and separation Turbulent Mixing and Chemical Reactions presents a balanced treatment of the connection between mixing and reaction It contains theoretical aspects experimental methods and expected results as well as worked examples to illustrate problem solving This book will be of interest to all scientists involved in chemical engineering physical chemistry and synthetic chemists in the fine chemical and pharmaceuticals industry

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