

QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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Ian Pickup



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Best Practices to Help At-Risk Learners Franklin Schargel,2014-05-22 This book presents 46 research based tools worksheets and resources which have been field tested at schools and dropout prevention programs across the country For teachers administrators counselors and special educators who work with at risk learners it helps you apply the Fifteen Strategies identified by The National Dropout Prevention Center and Network at Clemson University A set of indexes by grade level by job title and by individual strategy makes it easy for you to identify the tools worksheets and resources which will be of the greatest benefit to you and your students Examples include How To Encourage Students To Think About Staying In School How To Make Your School Family Friendly How to Identify High Performing At Risk Students What Can Parents and Teachers Do If an Adolescent Begins to Fail in School Reducing Special Education Dropouts and How to Get the Community Involved in Truancy and Dropout Prevention

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