

ATOMIC STRUCTURE AND THE PERIODIC TABLE  
CHAPTER 4 WORKSHEET

**PART A**

Given the following isotopes, determine the atomic number, the mass number, the number of protons, electrons and neutrons.

Isotope Symbol	Atomic Number	Mass Number	Protons	Electrons	Neutrons	Isotope Name
$^{131}_{53}\text{I}$	53	131	53	53	78	iodine - 131
$^{35}_{16}\text{S}$	16	35	16	16	19	sulphur - 35
$^4_2\text{He}$	2	4	2	2	2	helium - 4
$^{27}_{13}\text{Al}$	13	27	13	13	14	aluminium - 27
$^{81}_{36}\text{Kr}$	36	81	36	36	45	krypton - 81
$^{81}_{37}\text{Rb}$	37	81	37	37	44	rubidium - 81

**PART B**

Complete the following chart by writing the symbol for the isotope of the following elements. In addition, give the number of protons, electrons, mass number and atomic number and complete the element name.

Element Name	Neutrons	Protons	Electrons	Mass Number	Atomic Number	Isotope Symbol
Uranium - <u>237</u>	145	92	92	237	92	$^{237}_{92}\text{U}$
Chlorine - <u>45</u>	28	17	17	45	17	$^{45}_{17}\text{Cl}$
Oxygen - <u>17</u>	9	8	8	17	8	$^{17}_8\text{O}$
Boron - <u>11</u>	6	5	5	11	5	$^{11}_5\text{B}$
Beryllium - <u>9</u>	5	4	4	9	4	$^9_4\text{Be}$
Hydrogen - <u>2</u>	1	1	1	2	1	$^2_1\text{H}$
Carbon - <u>14</u>	8	6	6	14	6	$^{14}_6\text{C}$

# Physical Science Chapter 4 Answers

**R. M. Harbeck**



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