

Directions: Using the Reference Tables for Chemistry, locate the following information.

1. Name C_5H_{12} . **Pentane** Table P + R
2. Write the equation for the decay of Kr-85. $^{85}\text{Kr} \rightarrow \beta^- + ^{85}\text{Rb}$ Table N
3. Explain how you know that NaPO_4 is soluble in water but NiCrO_4 is not. **Na group 1/ all chromates insoluble** Table F
4. What is the definition of STP, and give the values? **Standard Temperature (0°C or 273K) and Pressure (1 atm or 101.3 kPa)** Table A
5. Name, and give the formulas of the strongest and weakest bases. **NaOH / NH_3** Table L
6. Name $C_2H_3O_2^-$ or CH_3COO^- **acetate ion** Table E
7. What is the solubility of sulfur dioxide at 40 degrees Celsius, in grams? **~5g** Table G
8. What is the freezing point of fluorine? **53 K** Table S
9. What are the units for the heat of fusion, and what do they mean? **J/g heat lost per gram of substance** Table B
10. What is the symbol for the mole? **mol** Table D
11. What is the vapor pressure of water at 75°C? **38 kPa** Table H
12. How much heat does it take to convert 20g of water to steam at 100°C? **45200 J (45.2 kJ)** I
13. What is the prefix for 1/1000 of a meter? **Milli-** Table C
14. What is the molecular formula of ammonia? **NH_3** Table L
15. What is the formula for the permanganate ion? **MnO_4^-** Table E
16. Name CH_3COOH **ethanoic acid (acetic acid)** Table R
17. Write the symbol for a positron? **β^+** Table O
18. What is the half-life of Pu-239? **2.4×10^4 years** Table N
19. Is the formation of water, from its elements endothermic or exothermic? **exothermic** Table I
20. What is the atomic mass of silver? **108 g/mol** P-Table
21. How much heat is released when LiBr dissolves in water? **-48.83 kJ** Table I
22. Give the names and formulas of the strongest and weakest acids. **HCl / CH_3COOH** Table K
23. What is the general formula for alkynes? What does it mean? **C_nH_{2n-2}** Table Q

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KJ Lindholm-Leary



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