

BALANCING REDOX REACTIONS

Oxidation Reduction



- During **oxidation**, oxidation number increases.
- During **reduction**, oxidation number decreases.

The reduced reactant is the oxidizing agent.

- H^+ gains electrons and changes to hydrogen gas by taking electrons from it. (its o.n. decreases from +1 to 0)

The oxidized reactant is the reducing agent.

- Zn loses electrons and changes to zinc ions. (its o.n. increases from 0 to +2)

**High School Chemistry
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R Barnett



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