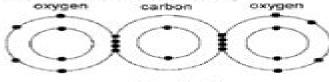


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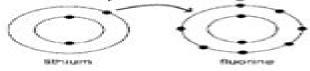
- I. MULTIPLE CHOICE INSTRUCTION: Encircle the letter of the correct answer. (1 POINT EACH)
- 1. It is a subatomic particle, symbol e- or
 - β-, with a negative elementary electric charge.
 - A. Neutron
 - 8. Electron
 - C. Protom
 - D. Attorni
- It is a subatomic particle, symbol p or p+, with a
 positive electric charge of +1e elementary
 charge and mass slightly less than that of a
 meutron.
 - A. Neutron
 - 8. Electron
 - C. Protom
 - Atom
- It is a subatomic particle, symbol n or nO, with no net electric charge and a mass slightly larger than that of a proton.
 - A. Neutron B. Electron
 - C. Proton
 - D. Atom
- 4. Is this an example of ionic bonding?



carbon dioxide

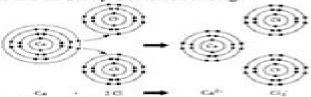
- A. Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- C. There is no chemical bonding present.
- D. This is just an orbital model.

5. Is this an example of ionic bonding?



littreum fluoride

- A. Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- There is no chemical bonding present.
- D. This is just an orbital model.
- In atomic physics and quantum chemistry, the electron configuration is the distribution of electrons of an atom or molecule in atomic or molecular orbitals.
 - A. Electron Configuration
 - Periodic Table of Elements
 - C. Orbital Models
 - D. Emergy Levels
- 7. Is this an example of ionic bonding?



- A. Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- There is no chemical bonding present.
- This is just an orbital model.
- It is a chemical rule of thumb that reflects observation that atoms of main-group elements tend to combine in such a way that each atom has eight electrons in its valence shell, giving it the same electronic configuration as a noble gas.
 - A. Octet Bulle
 - 8. Valence Electrons
 - C. Atomic Orbitals
 - D. Chemical Bonds
- It is the smallest constituent unit of ordinary matter that has the properties of a chemical element.
 - A. Nieutron
 - B. Electron
 - C. Proton
 - D. Atom
- 10. It is a type of chemical bond that involves the electrostatic attraction between oppositely

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G.M
Srimyvizhiy,M.Jagadeeswaran,T.Prabh
a,V.Lalitha,Ravi Patel,Harsh J.
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Arguing with Numbers James Wynn,G. Mitchell Reyes,2021-04-15 As discrete fields of inquiry rhetoric and mathematics have long been considered antithetical to each other That is if mathematics explains or describes the phenomena it studies with certainty persuasion is not needed This volume calls into question the view that mathematics is free of rhetoric Through nine studies of the intersections between these two disciplines Arguing with Numbers shows that mathematics is in fact deeply rhetorical Using rhetoric as a lens to analyze mathematically based arguments in public policy political and economic theory and even literature the essays in this volume reveal how mathematics influences the values and beliefs with which we assess the world and make decisions and how our worldviews influence the kinds of mathematical instruments we construct and accept In addition contributors examine how concepts of rhetoric such as analogy and visuality have been employed in mathematical and scientific reasoning including in the theorems of mathematical physicists and the geometrical diagramming of natural scientists Challenging academic orthodoxy these scholars reject a math equals truth reduction in favor of a more constructivist theory of mathematics as dynamic evolving and powerfully persuasive By bringing these disparate lines of inquiry into conversation with one another Arguing with Numbers provides inspiration to students established scholars and anyone inside or outside rhetorical studies who might be interested in exploring the intersections between the two disciplines In addition to the editors the contributors to this volume are Catherine Chaput Crystal Broch Colombini Nathan Crick Michael Dreher Jeanne Fahnestock Andrew C Jones Joseph Little and Edward Schiappa

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Physical Science Grade 11 Final Examination 2014 Scope**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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