

Chapter 5 Study Guide- Physical Science  
Test Monday, April 28<sup>th</sup>, 2014

1. What is a chemical bond? **force of attraction that holds elements together in order to form a unit**
2. Define chemical structure. **the arrangement of atoms in a substance.**
3. What is an ionic bond? **a chemical bond in which one atom loses an electron to form a positive ion and the other atom gains an electron to form a negative ion**
4. An ionic bond is between a   **METAL**   and a   **NONMETAL**  .
5. What is a covalent bond? **a chemical bond that involves sharing a pair of electrons between atoms in a molecule**
6. A covalent bond is between a   **nonmetal**   and a   **nonmetal**  .
7. What is a metallic bond? **a bond formed by the attraction between positively charged metal ions and the electrons around them.**
8. What are the polyatomic ions? **an ion made of two or more atoms.**
9. What is an ION? **a charged atom**
10. What is a CATION? **A positively charged atom**
11. What is an ANION? **A negatively charged atom**
12. Which part of an ionic compound should be written first? **Metal**

**Name the following ionic compounds:**

13.  $\text{FeI}_2$ - **Iron Iodide**
14.  $\text{MnF}_3$ - **Manganese Fluoride**
15.  $\text{CrCl}_2$ - **Chromium Chloride**
16.  $\text{CuS}$ - **Copper Sulfide**

**Name the following covalent compounds:**

17.  $\text{As}_2\text{O}_5$ - **diarsenic pentaoxide**
18.  $\text{SiI}_4$ - **silicon tetraiodide**
19.  $\text{P}_4\text{S}_3$ - **tertraphosphorus trisulfide**
20.  $\text{P}_4\text{O}_{10}$ - **tetraphosphorus decoxide**

**Write the equation for the following ionic compounds:**

21. Lithium Oxide-  **$\text{Li}_2\text{O}$**
22. Magnesium Nitride-  **$\text{Mg}_3\text{N}_2$**
23. Sodium Phosphide-  **$\text{Na}_3\text{P}$**
24. Calcium Bromide - **$\text{CaBr}_2$**

**Define/Explain the following**

25. What is a subscript? **number written below, tells you how many of each element is present in compound/molecule**
26. What is a superscript? **number written above the element symbol, this is how the oxidation number (charge) is displayed.**

# Physical Science Chemical Bond Study Guide

**JR Anderson**



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## **Table of Contents Physical Science Chemical Bond Study Guide**

1. Understanding the eBook Physical Science Chemical Bond Study Guide
  - The Rise of Digital Reading Physical Science Chemical Bond Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chemical Bond Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Chemical Bond Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chemical Bond Study Guide
  - Personalized Recommendations
  - Physical Science Chemical Bond Study Guide User Reviews and Ratings
  - Physical Science Chemical Bond Study Guide and Bestseller Lists
5. Accessing Physical Science Chemical Bond Study Guide Free and Paid eBooks

- Physical Science Chemical Bond Study Guide Public Domain eBooks
- Physical Science Chemical Bond Study Guide eBook Subscription Services
- Physical Science Chemical Bond Study Guide Budget-Friendly Options
- 6. Navigating Physical Science Chemical Bond Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science Chemical Bond Study Guide Compatibility with Devices
  - Physical Science Chemical Bond Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Chemical Bond Study Guide
  - Highlighting and Note-Taking Physical Science Chemical Bond Study Guide
  - Interactive Elements Physical Science Chemical Bond Study Guide
- 8. Staying Engaged with Physical Science Chemical Bond Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Chemical Bond Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Chemical Bond Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Chemical Bond Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemical Bond Study Guide
  - Setting Reading Goals Physical Science Chemical Bond Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemical Bond Study Guide
  - Fact-Checking eBook Content of Physical Science Chemical Bond Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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