#### Chapter 9

#### 9.1 Compounds and Chemical Change

- Atom- the smallest unit of an element that can exist alone or in combination with other elements.
- Compounds- formed when atoms are held together by an attractive force called a chemical bond
- Molecule- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- Chemical reaction

   a change in matter where different chemical substances are created by breaking and/or forming chemical bonds.
  - Examples: Photosynthesis
- Energy is absorbed -> produce new chemical substances -> more internal potential energy
- Energy is released when new chemical substances are produced -> less internal potential energy
- Chemical Equation: a way of describing what happens during a chemical reaction
  - Example: Photosynthesis equation-

Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

#### 9.2 Valence Electrons and Ions

- the number of electrons in the outermost orbital that usually determines the chemical properties of an atom
- <u>Valence electrons</u> the number of electrons in the outermost orbital that usually determines the properties of an atom
- Octet Rule after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions

#### 9.3 Chemical Bonds

- chemical bond-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
  - Ionic Bonds
  - Covalent Bonds
  - Metallic Bonds
- Fonic Bonds
   - type of linkage formed from the electrostatic attraction between oppositely
   charged ions in a chemical compound
- Covalent bonds: result when atoms achieve the noble gas electron structure by sharing
  electrons. They generally formed between the nonmetallic elements on the right side of
  the periodic table
- Metallic bonds: formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- . Heat formation- energy that is released during an ionic bond formation
- Boemula- describes what elements are in the compound and in what proportions
- Simple bond covalent bond
- Double bond- formed by two pairs of electrons that are shared by two atoms

# **Physical Science Chapter Vocabulary**

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## **Table of Contents Physical Science Chapter Vocabulary**

- 1. Understanding the eBook Physical Science Chapter Vocabulary
  - The Rise of Digital Reading Physical Science Chapter Vocabulary
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Chapter Vocabulary
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Chapter Vocabulary
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Chapter Vocabulary
  - Personalized Recommendations

- Physical Science Chapter Vocabulary User Reviews and Ratings
- Physical Science Chapter Vocabulary and Bestseller Lists
- 5. Accessing Physical Science Chapter Vocabulary Free and Paid eBooks
  - Physical Science Chapter Vocabulary Public Domain eBooks
  - Physical Science Chapter Vocabulary eBook Subscription Services
  - Physical Science Chapter Vocabulary Budget-Friendly Options
- 6. Navigating Physical Science Chapter Vocabulary eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Physical Science Chapter Vocabulary Compatibility with Devices
  - Physical Science Chapter Vocabulary Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Chapter Vocabulary
  - Highlighting and Note-Taking Physical Science Chapter Vocabulary
  - Interactive Elements Physical Science Chapter Vocabulary
- 8. Staying Engaged with Physical Science Chapter Vocabulary
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Chapter Vocabulary
- 9. Balancing eBooks and Physical Books Physical Science Chapter Vocabulary
  - ∘ Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Chapter Vocabulary
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter Vocabulary
  - Setting Reading Goals Physical Science Chapter Vocabulary
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter Vocabulary
  - Fact-Checking eBook Content of Physical Science Chapter Vocabulary

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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