

RUNESCAPE 1-99 RANGED GUIDE



Runescape Range 1 99 Guide

J Spring



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web 1 consider the reaction $\text{i}_2\text{o}_5 \text{ g } 5 \text{ co g } 5 \text{ co}_2 \text{ g i}_2 \text{ g a } 80 \text{ 0 grams of iodine v oxide i}_2\text{o}_5 \text{ reacts with } 28 \text{ 0 grams of carbon monoxide co}$ determine the mass of iodine i_2 which could be produced b if in the above situation only 0 160 moles of iodine i_2 was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key 4 $\text{ko}_2 \text{ s } 2 \text{ h } 2 \text{ o l } 4 \text{ koh aq } 3 \text{ o}_2 \text{ g}$ if a reaction vessel contains 0 mol ko_2 and 0 mol $\text{h}_2 \text{ o}$ what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H₂S \times 18 g H₂O / 1 mole H₂O = 149.4 g H₂O is your theoretical yield so percent yield = 137.1 g H₂O actual yield / 149.4 g H₂O \times 100 = 91.77% problems 6 & 8 are similar to problem 5

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web 1.csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = mass of product obtained / calculated mass of the product expected \times 100 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃

$$160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g Al}} \times \frac{2 \text{ mole Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 150.0 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{126.0 \text{ g}}{150.0 \text{ g}} \times 100 = 84.0\%$$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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