

Physical Science A

Final Exam Study Guide

- **Matter**
 - Compare and contrast physical and chemical properties, and physical and chemical changes
 - Compare and contrast properties and behaviors of atoms in each of the 3 states of matter
 - Convert between mass, volume, and density, and relate appropriate SI units for each variable
 - Predict how solutions of different densities would settle in a graduated cylinder
- **Atoms, Elements, and the Periodic Table**
 - Describe Antoine Lavoisier and Dmitri Mendeleev's contributions to the development of the modern periodic table
 - Define the term ion and be able to predict the most common ions formed for elements by looking at the periodic table
 - Identify three subatomic particles and compare their properties
 - Compare and contrast atoms in ground and excited states
 - Distinguish the atomic number of an element from the mass number of an isotope, and use these numbers to describe the structure of atoms
 - Describe the currently accepted model of an atom
 - Interpret the information provided about the atoms and elements in a balanced chemical reaction
 - Differentiate between elements that are good and poor conductors of heat and electricity
 - Understand the development and organization of the periodic table, especially in terms of group properties, atomic mass and valence electrons
 - Identify general properties of metals, nonmetals, and metalloids
 - Define the words malleable, brittle, volatile, and ductile
 - Explain how the electron cloud model represents the behavior and locations of electrons in atoms
 - Predict the reactivity of some elements based on their locations in the periodic table
 - Compare and contrast Bohr and Rutherford's atomic models
 - Distinguish the atomic number of an element from the mass number of an isotope, and use these numbers to describe the structure of atoms
- **Bonding**
 - List the monomers that make up polymers for different biological macromolecules
 - Compare and contrast ionic and covalent bonds
 - Understand the development and organization of the periodic table, especially in terms of group properties, atomic mass and valence electrons
 - Explain the role that valence electrons play in bonding
 - Explain the octet rule for bonding
 - Draw Lewis dot diagrams for elements, and use them to predict bonding between one or more atoms
 - Identify three subatomic particles and compare their properties
 - Compare and contrast atoms in ground and excited states
- **Chemical Reactions**
 - Describe factors that can be used to increase reaction rates
 - Classify chemical reactions as synthesis, decomposition, single-replacement, double-replacement, or combustion reactions
 - Predict the products of chemical reactions
 - Interpret chemical equations in terms of reactants, products, and conservation of mass
 - Balance chemical equations by manipulating coefficients
 - Calculate amounts of reactants or products by using molar mass, molar ratios, and balanced chemical equations
 - Apply the Law of Conservation of Mass to chemical reactions
 - Use signs to tell that a chemical reaction has occurred
 - Distinguish between endothermic and exothermic reactions

Note Taking Study Guide Physical Science

JR Anderson



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