

Chapter 12 – Stoichiometry

Note Taking Guide: Episode 801 Stoichiometry

-Study of the **QUANTITY** relationships in a **CHEMICAL REACTION**

-Based on **BALANCED** equations



1. Particles

2 atoms Mg

2 atoms O

2 atoms Mg

2 atoms O

2 atoms Mg

1 molecule O₂

2 molecule MgO

10 atoms Mg

5 molecule O₂

10 molecule MgO

2 (6.02 x 10²³)
atoms Mg

1 (6.02 x 10²³)
molecule O₂

2 (6.02 x 10²³)
molecule MgO

2. Moles

2 mol Mg

1 mol O₂

2 mol MgO

The **COEFFICIENTS** in a **BALANCED EQUATION** give the **MOLE RATIO** for the **SUBSTANCES** involved in the **REACTION**.

Lab: Episode 801 - Can You Find the Mole Ratio?

In this experiment, baking soda (NaHCO₃) reacts with hydrochloric acid (HCl) to produce sodium chloride, carbon dioxide, and water. You will determine the moles of reactant used and moles of product produced using both the lab data and stoichiometry and compare the two.

Data Table:

mass of test tube and baking soda	19.46 g
mass of empty test tube	19.21 g

Note Taking Guide Episode 801 Stoichiometry

Mark Steyvers



Note Taking Guide Episode 801 Stoichiometry:

Reviewing **Note Taking Guide Episode 801 Stoichiometry**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Note Taking Guide Episode 801 Stoichiometry**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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