

SALT ANALYSIS



SCHEME OF ANALYSIS

EXPERIMENT	OBSERVATION	INFERENCE
PRELIMINARY EXAMINATIONS :		
1) Noted the colour	Colourless	Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$
	Blue	Pre. of Cu^{2+}
	Brown	Pre. of Fe^{3+}
	Green	Pre. of Ni^{2+}
	Pale pink	Pre. of Mn^{2+}
	Pink	Pre. of Co^{2+}
2) Noted the solubility a) In water b) In dil.HCl	a) Soluble and no ppt with Na_2CO_3 b) Insoluble	a) Pre. of NH_4^+ b) Pre. of Pb^{2+}
3) Flame Test : A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.	a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating	a) Pre. of Cu^{2+} b) Pre. of Ca^{2+} c) Pre. of Sr^{2+} d) Pre. of Pb^{2+} e) Pre. of Ba^{2+}
4) Ash Test : A little of the salt is heated with con. HNO_3 and cobalt nitrate solution. A filter paper is dipped in it and burnt.	a) Green ash b) Blue tinted ash c) Pale pink ash	a) Pre. of Zn^{2+} b) Pre. of Al^{3+} c) Pre. of Mg^{2+}
SYSTEMATIC ANALYSIS OF ANIONS :		
1) $\text{dil.H}_2\text{SO}_4$ test : To the salt add $\text{dil.H}_2\text{SO}_4$, warm if required.	a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky. b) Smell of vinegar	a) Pre. of carbonates b) Pre. of acetate

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JL Elias



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Table of Contents Salt Analysis Manual For P G Course

1. Understanding the eBook Salt Analysis Manual For P G Course
 - The Rise of Digital Reading Salt Analysis Manual For P G Course
 - Advantages of eBooks Over Traditional Books
2. Identifying Salt Analysis Manual For P G Course
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Salt Analysis Manual For P G Course
 - User-Friendly Interface
4. Exploring eBook Recommendations from Salt Analysis Manual For P G Course
 - Personalized Recommendations
 - Salt Analysis Manual For P G Course User Reviews and Ratings
 - Salt Analysis Manual For P G Course and Bestseller Lists
5. Accessing Salt Analysis Manual For P G Course Free and Paid eBooks
 - Salt Analysis Manual For P G Course Public Domain eBooks
 - Salt Analysis Manual For P G Course eBook Subscription Services
 - Salt Analysis Manual For P G Course Budget-Friendly Options
6. Navigating Salt Analysis Manual For P G Course eBook Formats
 - ePub, PDF, MOBI, and More
 - Salt Analysis Manual For P G Course Compatibility with Devices
 - Salt Analysis Manual For P G Course Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Salt Analysis Manual For P G Course
 - Highlighting and Note-Taking Salt Analysis Manual For P G Course
 - Interactive Elements Salt Analysis Manual For P G Course
8. Staying Engaged with Salt Analysis Manual For P G Course

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Salt Analysis Manual For P G Course
- 9. Balancing eBooks and Physical Books Salt Analysis Manual For P G Course
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Salt Analysis Manual For P G Course
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Salt Analysis Manual For P G Course
 - Setting Reading Goals Salt Analysis Manual For P G Course
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Salt Analysis Manual For P G Course
 - Fact-Checking eBook Content of Salt Analysis Manual For P G Course
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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