

Calculating Average Rates of Reaction

You are required to show work for all calculations.

1. During the combustion of methane, CH_4 , shown by the reaction



The concentration of methane was measured at various time intervals and the following results were obtained:

Time (s)	$[\text{CH}_4]$ ($\text{mol} \cdot \text{l}^{-1}$)
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

Solution:

$$\begin{aligned}
 \text{rate} &= \frac{\Delta[\text{CH}_4]}{\Delta t} \\
 &= \frac{[0.60 - 2.4]\text{mol}}{(40 - 10)\text{s} \times \text{L}} \\
 &= \frac{-1.8\text{mol}}{30\text{s} \times \text{L}} \\
 &= -0.06 \frac{\text{mol}}{\text{L} \times \text{s}}
 \end{aligned}$$

Note: the answer is negative because we are calculating the **rate of loss of methane**.

Calculating Reaction Rates using Stoichiometry

Example:

In the following decomposition reaction,



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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'EduGorilla Prep Experts' text.

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McMahon,Rosemary McMahon,Bohdan Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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Table of Contents Reaction Rate Practice Problems

1. Understanding the eBook Reaction Rate Practice Problems
 - The Rise of Digital Reading Reaction Rate Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rate Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rate Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rate Practice Problems
 - Personalized Recommendations
 - Reaction Rate Practice Problems User Reviews and Ratings
 - Reaction Rate Practice Problems and Bestseller Lists
5. Accessing Reaction Rate Practice Problems Free and Paid eBooks
 - Reaction Rate Practice Problems Public Domain eBooks
 - Reaction Rate Practice Problems eBook Subscription Services
 - Reaction Rate Practice Problems Budget-Friendly Options
6. Navigating Reaction Rate Practice Problems eBook Formats

- ePub, PDF, MOBI, and More
- Reaction Rate Practice Problems Compatibility with Devices
- Reaction Rate Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rate Practice Problems
 - Highlighting and Note-Taking Reaction Rate Practice Problems
 - Interactive Elements Reaction Rate Practice Problems
- 8. Staying Engaged with Reaction Rate Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rate Practice Problems
- 9. Balancing eBooks and Physical Books Reaction Rate Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rate Practice Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rate Practice Problems
 - Setting Reading Goals Reaction Rate Practice Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rate Practice Problems
 - Fact-Checking eBook Content of Reaction Rate Practice Problems
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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