

# Reaction Rates

## A Model for Reaction Rates

In your textbook, read about expressing reaction rates and explaining reactions and their rates.

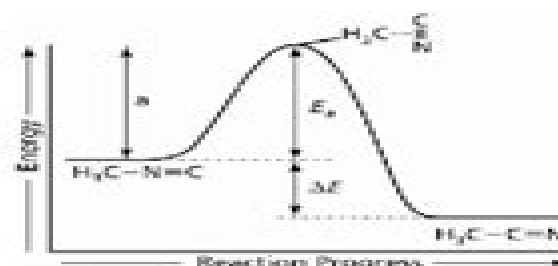
Use each of the terms below just once to complete the passage.

|                   |                   |          |
|-------------------|-------------------|----------|
| activated complex | activation energy | mol/L.s) |
| collision theory  | reaction rate     |          |

According to the (1) Collision theory, atoms, ions, and molecules must collide in order to react. Once formed, the (2) activated complex is a temporary, unstable arrangement of atoms that may then form products or may break apart to reform the reactants. Every chemical reaction requires energy, and the minimum amount of energy that reacting particles must have to form the activated complex is the (3) Activation energy. In a chemical reaction, the (4) reaction rate is the change in concentration of a reactant or product per unit time. It may be expressed using the units of (5) Mol/L.s).

Use the energy diagram for the rearrangement reaction of methyl isonitrile to acetonitrile to answer the following questions.

- What kind of reaction is represented by this diagram, endothermic or exothermic?  
exothermic
- What is the chemical structure identified at the top of the curve on the diagram?  
the activation complex
- What does the symbol  $E_a$  represent?  
the activation energy
- What does the symbol  $\Delta E$  represent?  
the net energy released from the exothermic reaction.



# Reaction Rates Study Guide Answers

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