

5 • Chemical Bonding

LEWIS STRUCTURES

Indicate the # of VALENCE electrons for each species. Write the correct Lewis electron dot structure for each.

F # of electrons = <u>7</u> $\cdot \ddot{\text{F}} \cdot$	O # of electrons = <u>6</u> $\begin{array}{c} \cdot \ddot{\text{O}} \cdot \\ \text{or} \cdot \ddot{\text{O}} \cdot \end{array}$	K # of electrons = <u>1</u> $\text{K} \cdot$	Al # of electrons = <u>3</u> $\cdot \ddot{\text{Al}} \cdot$
F⁻ <u>7+1</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{F}}:]^{-}$	O²⁻ <u>6+2</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{O}}:]^{2-}$	K⁺ <u>1-1</u> # of electrons = <u>0</u> $[\text{K}]^{+}$	Al³⁺ <u>3-3</u> # of electrons = <u>0</u> $[\text{Al}]^{3+}$
F₂ <u>7+7</u> # of electrons = <u>14</u> $\cdot \ddot{\text{F}} : \ddot{\text{F}} \cdot$	H₂ <u>1+1</u> # of electrons = <u>2</u> $\text{H} : \text{H}$	HF <u>1+7</u> # of electrons = <u>8</u> $\text{H} : \ddot{\text{F}} \cdot$	NH₃ <u>5+3</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} : \ddot{\text{N}} : \text{H} \\ \text{H} \end{array}$
CH₄ <u>4+4</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} \\ \\ \text{H} : \text{C} : \text{H} \\ \\ \text{H} \end{array}$	NF₃ <u>5+21</u> # of electrons = <u>26</u> $\begin{array}{c} \cdot \ddot{\text{F}} : \ddot{\text{N}} : \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} \cdot \end{array}$	SiF₄ <u>4+28</u> # of electrons = <u>32</u> $\begin{array}{c} \cdot \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} : \ddot{\text{Si}} : \ddot{\text{F}} \cdot \\ \\ \cdot \ddot{\text{F}} \cdot \end{array}$	C₂H₆ <u>8+6</u> # of electrons = <u>14</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} : \text{C} : \text{C} : \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
MgH₂ <u>2+2</u> # of electrons = <u>4</u> $\text{H} : \text{Mg} : \text{H}$	LiH <u>1+1</u> # of electrons = <u>2</u> $\text{Li} : \text{H}$	AlH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{Al} : \text{H} \\ \text{H} \end{array}$	BH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{B} : \text{H} \\ \text{H} \end{array}$

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