

EMPIRICAL AND MOLECULAR FORMULA

- 1) What is the empirical formula of a compound that contains 46.3 % lithium and 53.7% oxygen?
- 2) What is the empirical formula of a compound that contains 15.9 % boron and 84.1 % fluorine?
- 3) Phosphorus reacts with oxygen to give a compound that is 43.7% phosphorus and 56.4% oxygen. What is the empirical formula of the compound?
- 4) An inorganic salt is composed of 17.6% sodium, 39.7% chromium and 42.8% oxygen. What is the empirical formula of this salt?
- 5) Compound X contains 69.9% carbon, 6.86% hydrogen and 23.3% oxygen. Determine the empirical formula of X.
- 6) Oxalic acid has the empirical formula CHO_2 . Its molar mass is 90 g/mol. What is the molecular formula of oxalic acid?
- 7) The empirical formula of codeine is $\text{C}_{18}\text{H}_{21}\text{NO}_3$. If the molar mass of codeine is 299 g/mol, what is its molecular formula?
- 8) A compound's molar mass is 240.28 g/mol. Its percentage composition is 75.0% carbon, 5.05% hydrogen and 20.0% oxygen. What is the compound's molecular formula?
- 9) The wintergreen plant produces methyl salicylate, or oil of wintergreen. It can also be prepared easily in the laboratory. Methyl salicylate is 63.1% carbon, 5.31% hydrogen and 31.6% oxygen. Calculate the empirical formula of this compound.
- 10) An inorganic salt is made up of 38.8% calcium, 20.0% phosphorus, and 41.2% oxygen.
 - a) What is the empirical formula of the compound?
 - b) On further analysis, each formula unit of this salt is found to contain two phosphate ions. Predict the molecular formula of this salt.

Answers

- 1) Li_2O
- 2) BF_3
- 3) P_2O_5
- 4) $\text{Na}_2\text{Cr}_2\text{O}_7$
- 5) $\text{C}_{12}\text{H}_{14}\text{O}_3$
- 6) $\text{C}_2\text{H}_2\text{O}_4$
- 7) $\text{C}_{18}\text{H}_{21}\text{NO}_3$
- 8) $\text{C}_{15}\text{H}_{12}\text{O}_3$
- 9) $\text{C}_8\text{H}_8\text{O}_3$
- 10) a) $\text{Ca}_3\text{P}_2\text{O}_8$ b) $\text{Ca}_3(\text{PO}_4)_2$

Practice Empirical Formula Problems

Tobias Bleicker



Practice Empirical Formula Problems:

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Table of Contents Practice Empirical Formula Problems

1. Understanding the eBook Practice Empirical Formula Problems
 - The Rise of Digital Reading Practice Empirical Formula Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Empirical Formula Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Empirical Formula Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Empirical Formula Problems
 - Personalized Recommendations

- Practice Empirical Formula Problems User Reviews and Ratings
- Practice Empirical Formula Problems and Bestseller Lists
- 5. Accessing Practice Empirical Formula Problems Free and Paid eBooks
 - Practice Empirical Formula Problems Public Domain eBooks
 - Practice Empirical Formula Problems eBook Subscription Services
 - Practice Empirical Formula Problems Budget-Friendly Options
- 6. Navigating Practice Empirical Formula Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Empirical Formula Problems Compatibility with Devices
 - Practice Empirical Formula Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Empirical Formula Problems
 - Highlighting and Note-Taking Practice Empirical Formula Problems
 - Interactive Elements Practice Empirical Formula Problems
- 8. Staying Engaged with Practice Empirical Formula Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Empirical Formula Problems
- 9. Balancing eBooks and Physical Books Practice Empirical Formula Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Empirical Formula Problems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Empirical Formula Problems
 - Setting Reading Goals Practice Empirical Formula Problems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Empirical Formula Problems
 - Fact-Checking eBook Content of Practice Empirical Formula Problems

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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