

## AP Questions: Thermochemistry

1984 B

Substance	Standard Heat of Formation, $\Delta H_f^\circ$ in kJ/mol
C(s)	0.00
CO <sub>2</sub> (g)	-393.5
H <sub>2</sub> (g)	0.00
H <sub>2</sub> O(l)	-285.83
O <sub>2</sub> (g)	0.00
C <sub>3</sub> H <sub>7</sub> COOH(l)	?

The enthalpy change for the combustion of butyric acid at 25°C,  $\Delta H_{\text{comb}}$ , is -2,183.5 kilojoules per mole. The combustion reaction is



- From the above data, calculate the standard heat of formation,  $\Delta H_f^\circ$ , for butyric acid.
- Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

Substance	Enthalpy of Combustion, $\Delta H$ (Kilojoules/mole)
C(s)	-393.5
H <sub>2</sub> (g)	-285.8
C <sub>2</sub> H <sub>5</sub> OH(l)	-1366.7
H <sub>2</sub> O(l)	--

- Write a separate, balanced chemical equation for the combustion of each of the following: C(s), H<sub>2</sub>(g), and C<sub>2</sub>H<sub>5</sub>OH(l). Consider the only products to be CO<sub>2</sub> and/or H<sub>2</sub>O(l).
- In principle, ethanol can be prepared by the following reaction:



Calculate the standard enthalpy change,  $\Delta H$ , for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard molar enthalpy of neutralization of a strong acid by a strong base. Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- What equipment would be needed?
- What measurements should be taken?
- Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, C<sub>3</sub>H<sub>8</sub>, is a hydrocarbon that is commonly used as fuel for cooking.

- Write a balanced equation for the complete combustion of propane gas, which yields CO<sub>2</sub>(g) and H<sub>2</sub>O(l).
- Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O<sub>2</sub> by volume.
- The heat of combustion of propane is -2,220.1 kJ/mol. Calculate the heat of formation,  $\Delta H_f^\circ$ , of propane given that  $\Delta H_f^\circ$  of H<sub>2</sub>O(l) = -285.3 kJ/mol and  $\Delta H_f^\circ$  of CO<sub>2</sub>(g) = -393.5 kJ/mol.
- Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 kilograms of water (specific heat = 4.18 J/g°C), calculate the increase in temperature of water.

# Objectives Questions On Thermochemistry

**Dr. B. Kapila, S. K. Khanna**



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## **Table of Contents Objectives Questions On Thermochemistry**

1. Understanding the eBook Objectives Questions On Thermochemistry
  - The Rise of Digital Reading Objectives Questions On Thermochemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Objectives Questions On Thermochemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Objectives Questions On Thermochemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Objectives Questions On Thermochemistry



- Personalized Recommendations
  - Objectives Questions On Thermochemistry User Reviews and Ratings
  - Objectives Questions On Thermochemistry and Bestseller Lists
5. Accessing Objectives Questions On Thermochemistry Free and Paid eBooks
    - Objectives Questions On Thermochemistry Public Domain eBooks
    - Objectives Questions On Thermochemistry eBook Subscription Services
    - Objectives Questions On Thermochemistry Budget-Friendly Options
  6. Navigating Objectives Questions On Thermochemistry eBook Formats
    - ePub, PDF, MOBI, and More
    - Objectives Questions On Thermochemistry Compatibility with Devices
    - Objectives Questions On Thermochemistry Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Objectives Questions On Thermochemistry
    - Highlighting and Note-Taking Objectives Questions On Thermochemistry
    - Interactive Elements Objectives Questions On Thermochemistry
  8. Staying Engaged with Objectives Questions On Thermochemistry
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Objectives Questions On Thermochemistry
  9. Balancing eBooks and Physical Books Objectives Questions On Thermochemistry
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Objectives Questions On Thermochemistry
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Objectives Questions On Thermochemistry
    - Setting Reading Goals Objectives Questions On Thermochemistry
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Objectives Questions On Thermochemistry

- Fact-Checking eBook Content of Objectives Questions On Thermochemistry
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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