

Question	Expected Answers	Marks	Additional Guidance
3 a	measured pH > 1 OR $[H^+] < 0.1 \text{ (mol dm}^{-3}\text{)}$ ✓ $[H^+] = 10^{-0.4}$ ✓ $K_a = \frac{[H^+][CH_3CH_2COO^-]}{[CH_3CH_2COOH]}$ OR $\frac{[H^+]^2}{[CH_3CH_2COOH]}$ ✓ Calculate K_a from $\frac{[H^+]^2}{0.100}$ ✓	4	ALLOW C_2H_5 throughout question ALLOW $[H^+] < [CH_3CH_2COOH]$ OR $[H^+] < [HA]$ ALLOW measured pH is higher than expected ALLOW measured pH is not as acidic as expected ALLOW a quoted pH value or range > 1 and < 7 OR between 1 and 7 ALLOW $[H^+] = \text{antilog } -\text{pH}$ OR $[H^+] = \text{inverse log } -\text{pH}$ ALLOW $\frac{[H^+][A^-]}{[HA]}$ OR $\frac{[H^+]^2}{[HA]}$ IF K_a is NOT given and $K_a = \frac{[H^+]^2}{0.100}$ is shown, award mark for K_a also (I.e. $K_a = \frac{[H^+]^2}{0.100}$ is automatically awarded the last 2 marks)
b	Marks are for correctly calculated values. Working shows how values have been derived. $[H^+] = 10^{-13.46} = 3.47 \times 10^{-14} \text{ (mol dm}^{-3}\text{)}$ ✓ $[OH^-] = \frac{1.0 \times 10^{-14}}{3.47 \times 10^{-14}} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓	2	ALLOW $3.467368505 \times 10^{-14}$ and correct rounding to 3.5×10^{-14} ALLOW 0.28840315 and correct rounding to 0.29, i.e. ALLOW 0.288 ALLOW alternative approach using pOH: pOH = $14 - 13.46 = 0.54$ ✓ $[OH^-] = 10^{-0.54} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓ Correct answer gets BOTH marks

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Daniela Niemeyer



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